



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/52

Paper 5 Practical Test

March 2017

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the March 2017 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.

© IGCSE is a registered trademark.

This syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **4** printed pages.



Question	Answer	Marks
1(b)	initial and final readings completed correctly	1
	differences completed correctly	1
	all readings to 1 decimal place	1
	results comparable to the Supervisor's results	1
1(c)	blue to colourless	1
1(d)	neutralisation	1
1(e)(i)	solution O	1
	greater volume of acid was used in the titration	1
1(e)(ii)	five times as concentrated	1
1(f)	half volume / value from table for Experiment 2	1
	unit: cm ³	1
1(g)	effect on volume: no effect	1
	reason: temperature would only affect the rate	1
1(h)(i)	use a pipette / burette	1
1(h)(ii)	repeat experiments (and compare / average)	1

Question	Answer	Marks
1(i)	M1 fair test to equal volumes of each sodium hydroxide solution / solutions O and P add an equal volume / measured volumes of aqueous calcium chloride	1
	M2 dependent variable measured measure mass / height of precipitate formed / volume of calcium chloride used	1
	M3 conclusion the more concentrated sodium hydroxide solution would form the most precipitate (mass / height) / would require a smaller volume of calcium chloride	1

Question	Answer	Marks
2(a)	<u>white</u> (wet) (crystals)	1
2(b)(i)	white	1
	precipitate	1
2(b)(ii)	precipitate dissolves / clears / is soluble	1
2(c)(i)	white	1
	precipitate	1
2(c)(ii)	precipitate dissolves	1
2(d)	cream precipitate	1
2(e)	zinc	1
	bromide	1
2(f)	yellow	1
2(g)	no reaction / no change / precipitate	1

Question	Answer	Marks
2(h)	yellow	1
	precipitate	1
2(i)	sodium	1
	iodide	1

Question	Answer	Marks
3(a)	any 4 from: M1 measure initial temperature of (solid) ammonium chloride / barium hydroxide M2 add barium hydroxide / ammonium chloride / other solid AND mix / stir M3 use a thermometer M4 measure the temperature of the mixture / final temperature M5 temperature decreases / test-tube feels cold	4
3(b)	M1 add (aqueous) sodium hydroxide (and warm)	1
	M2 gas produced turns (red) litmus blue	1