



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0620/42

February/March 2024

1 hour 15 minutes

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

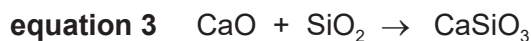
- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

IB24 03_0620_42/6RP
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[Turn over

- 1 Iron ore contains iron(III) oxide, Fe_2O_3 . A blast furnace is used to extract iron from Fe_2O_3 .

Equations for some of the reactions in the blast furnace are shown.



- (a) **Equation 1** shows the combustion of carbon in the blast furnace.

- (i) Name the substance which provides the carbon for this reaction.

..... [1]

- (ii) State the purpose of the combustion of carbon in the blast furnace.

..... [1]

- (b) Iron(III) oxide, Fe_2O_3 , in iron ore is converted to iron when it reacts with carbon monoxide, CO , in the blast furnace.

- (i) Calculate the percentage by mass of iron in iron(III) oxide, Fe_2O_3 .

percentage =% [2]

- (ii) State the name of the iron ore which consists mainly of iron(III) oxide.

..... [1]

- (iii) Describe how carbon monoxide is formed in the blast furnace.

..... [1]

- (iv) Write the symbol equation to show the reaction that occurs when iron(III) oxide is converted to iron in the blast furnace.

..... [2]

- (v) Name the chemical process which happens to iron when iron(III) oxide is converted to iron in the blast furnace.

..... [1]

- (c) State the type of reaction shown by **equation 2**.

..... [1]

(d) (i) Explain why the reaction in **equation 3** can be described as an acid–base reaction.

.....

.....

..... [2]

(ii) State:

- the chemical name of SiO_2

.....

- the common name given to CaSiO_3 when it is formed in the blast furnace.

..... [2]

(e) Aluminium **cannot** be extracted from its ore using a blast furnace.

(i) State why aluminium is **not** extracted from its ore using a blast furnace.

..... [1]

(ii) Name the process used to extract aluminium from its ore.

..... [1]

(f) Both iron(III) oxide and aluminium oxide contain metal ions with a 3+ charge.

(i) Write the electronic configuration of an Al^{3+} ion.

..... [1]

(ii) Deduce the number of protons and electrons in an Fe^{3+} ion.

protons	electrons

[2]

[Total: 19]

- 2 The elements in Group VII of the Periodic Table are known as the halogens. Halogens can form halide ions.

(a) Identify the halogen with the lowest density at r.t.p. (room temperature and pressure).

..... [1]

(b) State the appearance of bromine at r.t.p.

..... [1]

(c) Use the Periodic Table to:

- give the symbol of the halogen with the highest atomic number

.....

- deduce the number of occupied electron shells in an atom of this element.

.....

[2]

(d) Bromine molecules have covalent bonding.

(i) State what is meant by the term covalent bond.

.....

..... [2]

(ii) Name **one** halide ion which bromine molecules can displace.

..... [1]

(iii) Explain why bromine can displace the halide ion in (d)(ii).

..... [1]

(e) Name a halide compound which can be used to detect the presence of water.

..... [2]

- (f) Calcium chloride is an ionic compound.

Complete the dot-and-cross diagram in Fig. 2.1 for the ions in calcium chloride.

Give the charges on each of the ions.

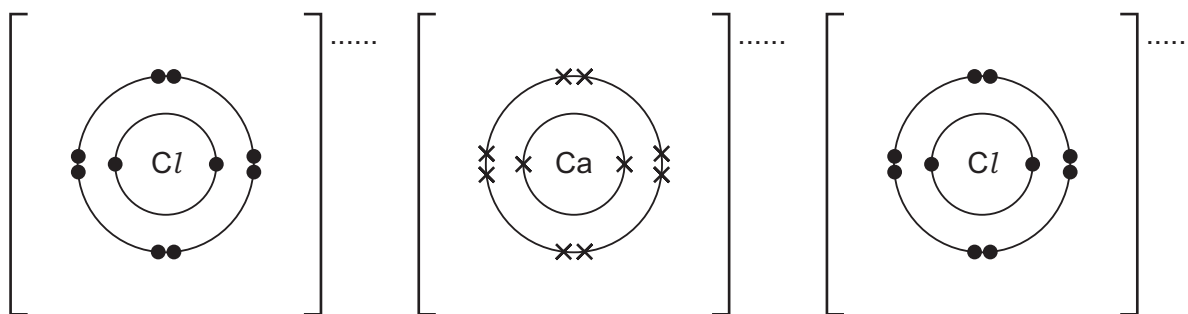


Fig. 2.1

[3]

- (g) Aqueous lead(II) ions are added to aqueous chloride ions. A white precipitate of insoluble lead(II) chloride, PbCl_2 , is formed.

- (i) Name a lead(II) compound which can be used in this reaction.

..... [1]

- (ii) Write the ionic equation for this reaction. Include state symbols.

.....
 [3]

- (iii) Name one **other** insoluble chloride.

..... [1]

[Total: 18]

- 3 This question is about acids, bases and alkalis.

Table 3.1 shows the pH values of some substances.

Table 3.1

substance	pH
NaOH(aq)	14
Ca(OH) ₂ (aq)	10
H ₂ O(l)	7
CH ₃ COOH(aq)	4
HNO ₃ (aq)	1

- (a) Define the term base.

..... [1]

- (b) State what is meant by the term alkali.

..... [1]

- (c) Thymolphthalein is an indicator.

State the colour of thymolphthalein in:

- NaOH(aq)
- CH₃COOH(aq).

[2]

- (d) (i) Use the information in Table 3.1 to identify the substance with the highest concentration of H⁺(aq) ions.

Explain your answer.

substance

explanation

[2]

- (ii) Name an indicator which can be used to identify the substance with the highest concentration of H⁺(aq) ions.

..... [1]

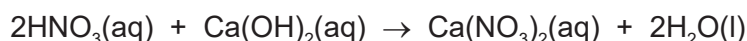
- (e) Complete the equation to show the dissociation of ethanoic acid, CH_3COOH , in aqueous solution.

$\text{CH}_3\text{COOH}(\text{aq})$ [3]

- (f) Write the **ionic** equation which represents a neutralisation reaction between any acid and any alkali.

..... [1]

- (g) Dilute nitric acid, $\text{HNO}_3(\text{aq})$, reacts with aqueous calcium hydroxide, $\text{Ca}(\text{OH})_2(\text{aq})$, as shown.



20.0 cm^3 of 0.0150 mol/dm^3 $\text{Ca}(\text{OH})_2(\text{aq})$ reacts with 25.0 cm^3 of $\text{HNO}_3(\text{aq})$.

Calculate the concentration of $\text{HNO}_3(\text{aq})$ in g/dm^3 .

Use the following steps.

- Calculate the number of moles of $\text{Ca}(\text{OH})_2(\text{aq})$ used.

..... mol

- Determine the number of moles of $\text{HNO}_3(\text{aq})$ which react with the $\text{Ca}(\text{OH})_2(\text{aq})$.

..... mol

- Calculate the concentration of $\text{HNO}_3(\text{aq})$ in mol/dm^3 .

..... mol/dm^3

- Calculate the concentration of $\text{HNO}_3(\text{aq})$ in g/dm^3 .

..... g/dm^3
[5]

[Total: 16]

- 4 The equation for the reaction between methanoic acid and ethanol in the presence of a catalyst can be represented as shown.



X represents the ester formed.

- (a) (i) In the equation, methanoic acid is represented by the formula HCOOH.

Name this type of formula.

..... [1]

- (ii) Write the empirical formula of methanoic acid.

..... [1]

- (b) Name and draw the displayed formula of ester **X**.

name

displayed formula

[3]

- (c) The reaction is reversible and reaches an equilibrium within a closed system.

- (i) State what is meant by the term closed system.

..... [1]

- (ii) State **two** characteristics of an equilibrium.

1

.....

2

.....

[2]

- (iii) Complete Table 4.1 to show the effect, if any, on the concentration of **X** at equilibrium for each change of condition.

Table 4.1

change of condition	effect on the concentration of X at equilibrium
temperature is decreased	
concentration of HCOOH is decreased	
concentrations of both HCOOH and CH ₃ CH ₂ OH are decreased	
the catalyst is removed	

[4]

[Total: 12]

5 Butane and but-1-ene are colourless gases at room temperature and pressure.

(a) Suggest why but-1-ene diffuses quicker than butane.

..... [1]

(b) Identify the products formed when butane undergoes complete combustion.

..... [1]

(c) One molecule of butane reacts with one molecule of chlorine in the presence of ultraviolet light. During the reaction, one hydrogen atom in butane is replaced by one chlorine atom.

(i) Name the type of reaction which needs ultraviolet light.

..... [1]

(ii) State the purpose of ultraviolet light during this reaction.

..... [1]

(iii) Name the type of reaction which takes place when one atom of chlorine replaces one atom of hydrogen.

..... [1]

(iv) Determine how many different structural isomers can form during this reaction.

..... [1]

(d) When but-1-ene reacts with steam, **two** possible products form.

(i) Identify the type of catalyst which is used in this reaction.

..... [1]

(ii) Name and draw the displayed formulae of the **two** possible products.

product 1	product 2
name	name
displayed formula	displayed formula

[4]

(e) But-1-ene undergoes polymerisation.

(i) State the type of polymerisation but-1-ene undergoes.

..... [1]

(ii) Draw part of the polymer molecule to show **three** repeat units.

[3]

[Total: 15]

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The Periodic Table of Elements

Group																			
I	II											III	IV	V	VI	VII	VIII		
		1 H hydrogen 1																	
		2 He helium 4																	
3	Li lithium 7	4	Be beryllium 9	atomic number atomic symbol name relative atomic mass										5	6	7	8	9	10
11	Na sodium 23	12	Mg magnesium 24											13	14	15	16	17	18
														Al aluminium 27	Si silicon 28	P phosphorus 31	S sulfur 32	Cl chlorine 35.5	Ar argon 40
																	</		