



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0620/31

May/June 2022

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

- 1 (a) A list of symbols and formulae is shown.

CaO
CH₄
C₂H₄
C₂H₆
Cl⁻
Cu²⁺
H₂
He
K⁺
N₂
Na⁺
SO₂

Answer the following questions using these symbols or formulae.
Each symbol or formula may be used once, more than once or not at all.

State which symbol or formula represents:

- (i) a compound produced by the thermal decomposition of calcium carbonate

..... [1]

- (ii) a positive ion that gives a blue-green colour in a flame test

..... [1]

- (iii) an element used as a fuel

..... [1]

- (iv) the monomer used to produce poly(ethene)

..... [1]

- (v) an ion formed when an atom gains an electron.

..... [1]

- (b) Complete the table to show the relative charges of a proton, a neutron and an electron.

type of particle	relative charge
proton	+1
neutron	
electron	

[2]

- (c) Choose the two correct statements about nitrogen.
Tick (✓) **two** boxes.

Nitrogen molecules are monoatomic.

☐

All nitrogen atoms have seven protons.

☐

Nitrogen atoms cannot be split into simpler substances by chemical means.

☐

All nitrogen atoms have 14 neutrons.

☐

Nitrogen is 21% of clean, dry air.

☐

[2]

[Total: 9]

- 2 The table shows the masses of some ions in a 1000 cm^3 sample of toothpaste.

name of ion	formula of ion	mass of ion in 1000 cm^3 of toothpaste/g
	NH_4^+	0.2
calcium	Ca^{2+}	1.2
	Cl^-	0.9
fluoride	F^-	1.4
magnesium	Mg^{2+}	2.0
phosphate	PO_4^{3-}	24.4
sodium	Na^+	28.1
sulfate	SO_4^{2-}	9.2
tin(II)	Sn^{2+}	0.2
zinc	Zn^{2+}	0.1

- (a) Answer these questions using only the information in the table.

- (i) State which negative ion has the lowest mass in 1000 cm^3 of toothpaste.

..... [1]

- (ii) Name the compound that contains NH_4^+ and Cl^- ions.

..... [1]

- (iii) Calculate the mass of phosphate ions in 250 cm^3 of toothpaste.

mass = g [1]

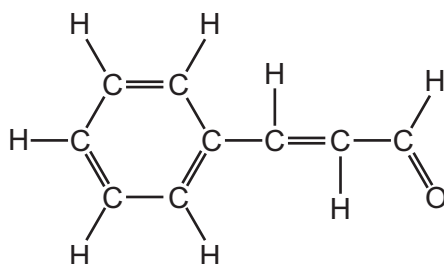
- (b) Describe a test for sulfate ions.

test

observations

[2]

- (c) Toothpaste also contains cinnamal.
The structure of cinnamal is shown.



Deduce the formula of cinnamal to show the number of atoms of carbon, hydrogen and oxygen.

..... [1]

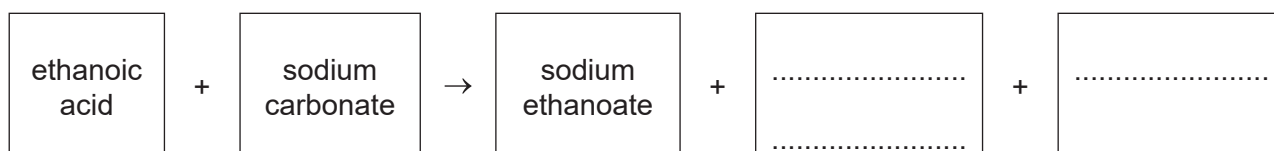
- (d) Cinnamal can be oxidised to a carboxylic acid.

- (i) Draw the structure of a carboxylic acid functional group to show all of the atoms and all of the bonds.

[1]

- (ii) Ethanoic acid is a carboxylic acid.
Ethanoic acid reacts like a typical acid.

Complete the word equation for the reaction of ethanoic acid with sodium carbonate.



[2]

[Total: 9]

3 This question is about halogens and halogen compounds.

- (a) Deduce the number of electrons, neutrons and protons in one atom of the isotope of chlorine shown.



number of electrons

number of neutrons

number of protons

[3]

- (b) State why chlorine is used in water treatment.

..... [1]

- (c) Aqueous chlorine reacts with aqueous potassium iodide.

- (i) Complete the chemical equation for this reaction.



- (ii) Explain in terms of the reactivity of the halogens why aqueous iodine does **not** react with aqueous potassium chloride.

.....

..... [1]

(d) The table shows some properties of four halogens.

halogen	melting point /°C	boiling point /°C	density of liquid at boiling point in g/cm ³
fluorine	−220		1.51
chlorine	−101	−35	
bromine	−7	59	3.12
iodine	114	184	4.93

(i) Complete the table by predicting:

- the boiling point of fluorine
- the density of liquid chlorine at its boiling point.

[2]

(ii) Predict the physical state of chlorine at −105°C.
Give a reason for your answer.

.....

..... [2]

[Total: 11]

4 This question is about acids and bases.

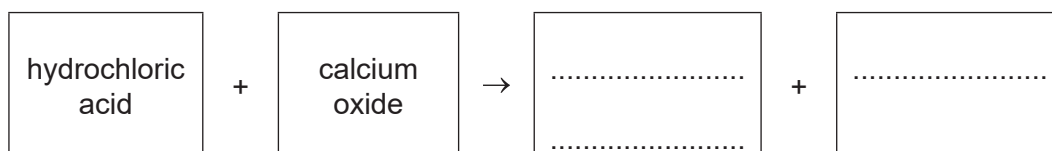
(a) Describe the colour of methyl orange in acidic and alkaline solutions.

in acidic solution

in alkaline solution

[2]

(b) Complete the word equation for the reaction of hydrochloric acid with calcium oxide.



[2]

(c) Calcium oxide is lime.

Give **one** use of lime.

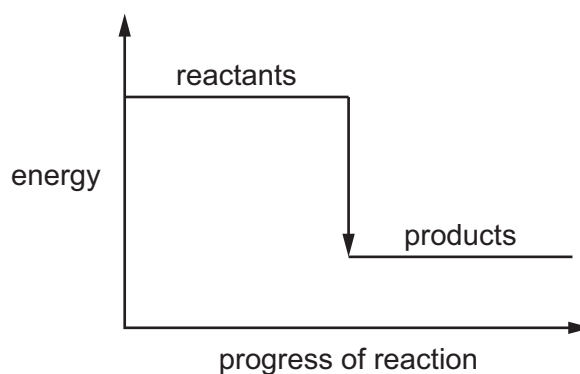
..... [1]

(d) The reaction of hydrochloric acid with calcium oxide is exothermic.

(i) State the meaning of the term *exothermic*.

..... [1]

(ii) The energy level diagram for the reaction of hydrochloric acid with calcium oxide is shown.



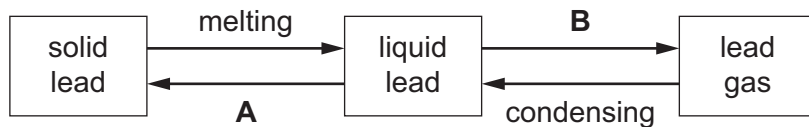
Explain how the energy level diagram shows that this reaction is exothermic.

.....
..... [1]

[Total: 7]

5 This question is about Group IV elements and their compounds.

(a) The changes of state of lead are shown.



Name the changes of state represented by **A** and **B**.

A

B

[2]

(b) Use the kinetic particle model to describe the differences between liquid lead and lead gas in terms of:

- the separation of the particles

.....

.....

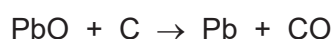
- the motion of the particles.

.....

.....

[4]

(c) Lead is extracted from lead(II) oxide by heating with carbon.



Describe how this equation shows that lead(II) oxide is reduced.

.....

..... [1]

(d) Lead is a pollutant of the air.

(i) State **one** source of lead in the air.

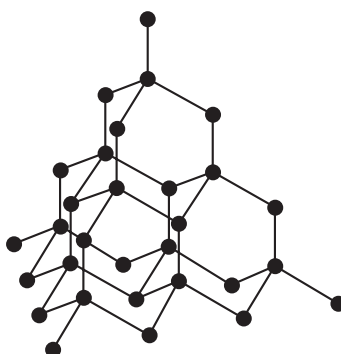
..... [1]

(ii) State **one** adverse effect of lead on health.

..... [1]

(e) Diamond is a form of carbon.

The structure of diamond is shown.



(i) Choose the word which best describes the structure of diamond.

Draw a circle around your chosen answer.

giant ionic metallic simple [1]

(ii) Name the type of bonding in diamond.

..... [1]

(iii) Give **one** use of diamond.

..... [1]

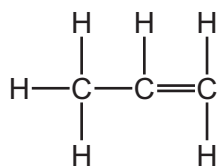
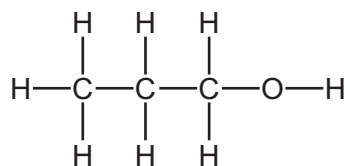
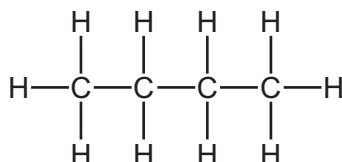
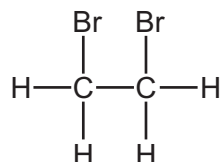
(iv) Deduce the electronic structure of carbon.

Use the Periodic Table to help you.

..... [1]

[Total: 13]

- 6 (a) The structures of four organic compounds, **C**, **D**, **E** and **F**, are shown.

C**D****E****F**

Answer the following questions about these compounds.
Each compound may be used once, more than once or not at all.

State which compound, **C**, **D**, **E** or **F**:

- (i) decolourises aqueous bromine [1]
- (ii) is an alcohol [1]
- (iii) is unsaturated [1]
- (iv) is in the same homologous series as ethane. [1]
- (b) Petroleum is a mixture of hydrocarbons which can be separated into fractions with different boiling points.

Name the method used to separate these fractions.

..... [1]

- (c) Complete the table to show the name and uses of some petroleum fractions.

name of fraction	use of fraction
refinery gas	
gasoline	
	waxes and polishes

[3]

(d) Some hydrocarbons are formed by the process of cracking.

(i) State the meaning of the term *cracking*.

.....

.....

..... [2]

(ii) Describe the conditions needed for cracking.

.....

..... [2]

[Total: 12]

7 This question is about zinc and compounds of zinc.

(a) Zinc is a metal.

Give **three** physical properties of metals.

1

2

3

[3]

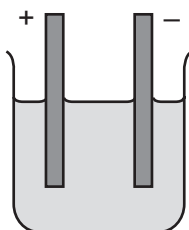
(b) Zinc reacts with phosphorus to form zinc phosphide, Zn_3P_2 .

Complete the equation for this reaction.



[2]

(c) Molten zinc chloride is electrolysed.
The incomplete apparatus is shown.



- (i) Complete the diagram by:
- completing the circuit to show the wires and power pack
 - labelling the anode.

[2]

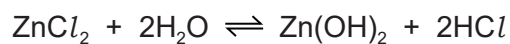
(ii) Name the products formed at each electrode.

positive electrode

negative electrode

[2]

(d) Zinc chloride reacts with water as shown. The solution formed is acidic.



(i) State the meaning of the symbol $\rightleftharpoons$.

..... [1]

(ii) Choose the pH value which is acidic.

Draw a circle around your chosen answer.

pH 3 pH 7 pH 9 pH 14 [1]

(e) A compound of zinc has the formula $\text{ZnC}_4\text{H}_{10}$.

Complete the table to calculate the relative molecular mass of $\text{ZnC}_4\text{H}_{10}$.

atom	number of atoms	relative atomic mass	
zinc	1	65	$1 \times 65 = 65$
carbon		12	
hydrogen		1	

relative molecular mass = [2]

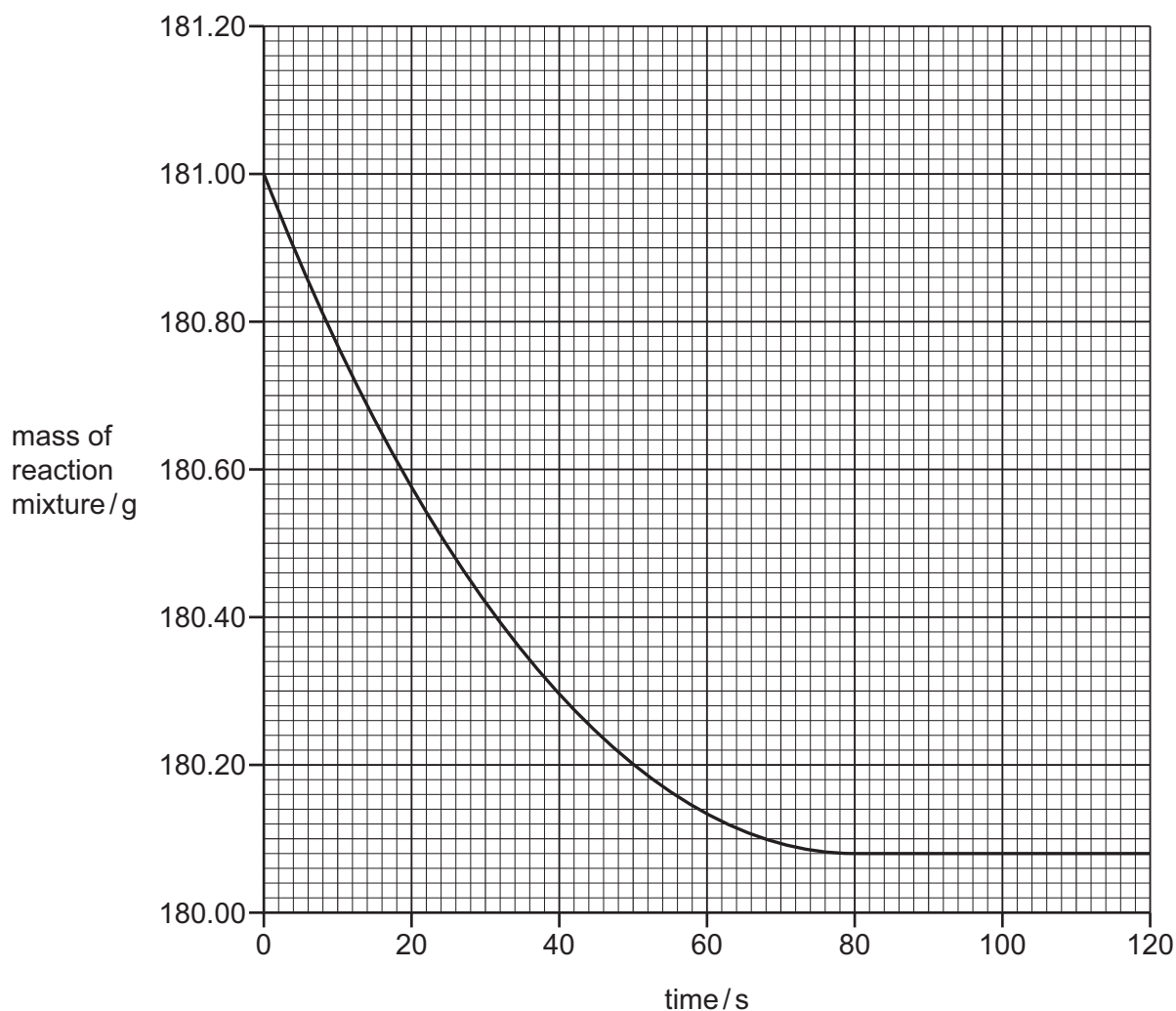
[Total: 13]

- 8 A student investigates the reaction of large pieces of copper(II) carbonate with dilute hydrochloric acid. The hydrochloric acid is in excess.



The rate of reaction is found by measuring the mass of the reaction mixture as time increases.

The results are shown on the graph.



- (a) Deduce the mass of the reaction mixture at 30 s.

mass = g [1]

- (b) The experiment is repeated using smaller pieces of copper(II) carbonate.

All other conditions stay the same.

Draw a line **on the grid** to show how the mass of the reaction mixture changes as time increases. [2]

- (c) Describe the effect each of the following has on the rate of reaction of copper(II) carbonate with dilute hydrochloric acid.

All other conditions stay the same.

- The reaction is carried out in the presence of a catalyst.

.....

- The reaction is carried out using a lower concentration of hydrochloric acid.

.....

[2]

- (d) When 0.2g of copper(II) carbonate is used, 38cm³ of carbon dioxide gas is produced.

Calculate the volume of carbon dioxide gas produced when 0.50g of copper(II) carbonate is used.

volume of carbon dioxide gas = cm³ [1]

[Total: 6]

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The Periodic Table of Elements

Group																								
I	II											III	IV	V	VI	VII	VIII							
1Hhydrogen1 atomic number atomic symbol name relative atomic mass																								
3Li lithium 7	4Be beryllium 9																							
11Na sodium 23	12Mg magnesium 24																							
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84							
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131							
55Cs caesium 133	56Ba barium 137	57–71 lanthanoids		72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	86Rn radon —						
87Fr francium —	88Ra radium —	89–103 actinoids		104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —		114Fl flerovium —		116Lv livermorium —								
lanthanoids																								
57La lanthanum 139	58Ce cerium 140	59Pr praseodymium 141	60Nd neodymium 144	61Pm promethium —	62Sm samarium 150	63Eu europium 152	64Gd gadolinium 157	65Tb terbium 159	66Dy dysprosium 163	67Ho holmium 165	68Er erbium 167	69Tm thulium 169	70Yb ytterbium 173	71Lu lutetium 175										
actinoids																								
89Ac actinium —	90Th thorium 232	91Pa protactinium 231	92U uranium 238	93Np neptunium —	94Pu plutonium —	95Am americium —	96Cm curium —	97Bk berkelium —	98Cf californium —	99Es einsteinium —	100Fm fermium —	101Md mendelevium —	102No nobelium —	103Lr lawrencium —										