



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0620/33

October/November 2023

1 hour 15 minutes

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.



1 A list of substances is shown.

ammonia
calcium oxide
carbon monoxide
cobalt(II) chloride
ethane
ethanol
ethene
oxygen
potassium oxide
sodium sulfate
sulfuric acid
water

Answer the following questions using only the substances from the list.
Each substance may be used once, more than once or not at all.

Give the name of the substance that:

(a) is a product of photosynthesis

..... [1]

(b) is a member of the alkene homologous series

..... [1]

(c) has an ion with a charge of 1–

..... [1]

(d) is used to remove sulfur dioxide in flue gas desulfurisation

..... [1]

(e) is the product formed in a hydrogen–oxygen fuel cell

..... [1]

(f) is used to test for water.

..... [1]

[Total: 6]

2 Hydrocarbons are compounds of carbon and hydrogen.

(a) State the meaning of the term compound.

.....
 [2]

(b) Fig. 2.1 shows a fractionating column for separating petroleum into different hydrocarbon fractions.

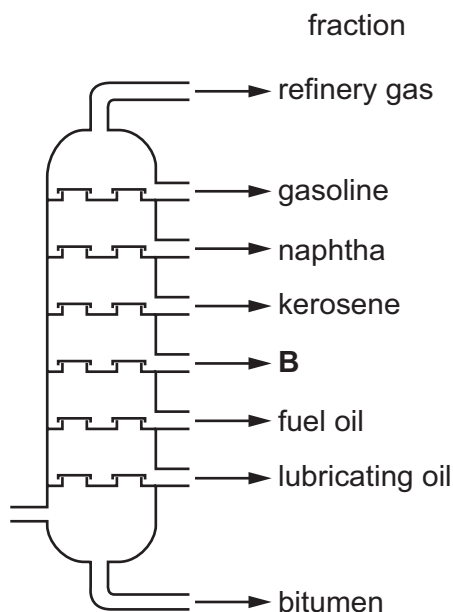


Fig. 2.1

(i) On Fig. 2.1, draw an **X** inside the column to show where the hydrocarbon with the highest boiling point collects. [1]

(ii) Name the fraction labelled **B** in Fig. 2.1.

..... [1]

(iii) State the name of the fraction which has hydrocarbons with the shortest chain length.

..... [1]

(iv) State **one** use of the naphtha fraction.

..... [1]

[Total: 6]

- 3 (a) Table 3.1 shows the average concentrations, in $\text{ng} / 1000 \text{cm}^3$, of air pollutants in four different years.

Table 3.1

year	concentration of air pollutant in $\text{ng} / 1000 \text{cm}^3$				
	carbon monoxide	hydrocarbons	oxides of nitrogen	particulates	sulfur dioxide
2019	5.3	22.0	15.6	19.0	20.0
2020	4.1	13.5	14.8	20.1	18.2
2021	5.8	14.8	22.7	23.5	16.2
2022	2.6	18.0	10.9	26.2	14.0

- (i) Name the pollutant which has the highest concentration in 2019.

..... [1]

- (ii) Name the pollutant that shows a continuous decrease in concentration from 2019 to 2022.

..... [1]

- (iii) Calculate the average mass, in ng, of hydrocarbons in a 200cm^3 sample of polluted air in 2019.

mass = ng [1]

- (b) (i) State **one** source of oxides of nitrogen in the air.

..... [1]

- (ii) Oxides of nitrogen contribute to acid rain.

Give one **other** effect of oxides of nitrogen in the air.

..... [1]

- (iii) Unpolluted water has a neutral pH.

Choose from the list the pH value of a neutral substance.

Draw a circle around your chosen answer.

pH 1 pH 6 pH 7 pH 14

[1]

(c) Nitrogen dioxide is an acidic oxide.

Choose an oxide from the list which is also an acidic oxide.

Tick (✓) **one** box.

copper(II) oxide	☐
magnesium oxide	☐
phosphorus(V) oxide	☐
sodium oxide	☐

[1]

(d) Sulfur dioxide reacts with oxygen to produce sulfur trioxide.

(i) Complete the symbol equation for this reaction.



(ii) State the meaning of the symbol $\rightleftharpoons$.

..... [1]

(iii) Sulfur trioxide reacts with calcium oxide to produce calcium sulfate.

Describe a test for sulfate ions.

test

observations

[2]

[Total: 12]

4 Nitrogen is a gas at room temperature.

(a) State **two** general properties of a gas.

- 1
-
- 2
-
- [2]

(b) Fig. 4.1 shows the physical states of nitrogen.

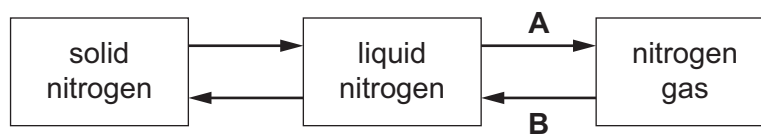


Fig. 4.1

Name the changes of physical states **A** and **B**.

- A**
- B**
- [2]

(c) Describe solid nitrogen and nitrogen gas in terms of the arrangement and separation of the particles.

solid nitrogen

arrangement

.....

separation

.....

nitrogen gas

arrangement

.....

separation

.....

[4]

- (d) A sealed gas syringe contains 80 cm^3 of nitrogen gas.

State how increasing the pressure affects the volume of nitrogen gas in the gas syringe when the temperature remains constant.

..... [1]

[Total: 9]

5 This question is about metals.

(a) Table 5.1 shows some properties of the Group I metals.

Table 5.1

metal	melting point in °C	boiling point in °C	atomic volume in cm ³ /mol	observations on reaction with water
lithium	181	1342	12.9	bubbles form slowly but no flame
sodium	98	883	23.7	
potassium	63	760		bubbles form very rapidly and flame seen
rubidium		686	55.8	explodes

Use the information in Table 5.1 to predict:

- (i) the melting point of rubidium [1]
- (ii) the atomic volume of potassium [1]
- (iii) the observations when sodium reacts with water [1]
- [1]
- (iv) the physical state of sodium at 1300 °C. Give a reason for your answer.
- physical state
- reason
- [2]

(b) Iron is extracted in a blast furnace by reduction of iron(III) oxide.

(i) In the first step, carbon burns in air to form carbon dioxide.

State the percentage of oxygen in clean, dry air.

..... [1]

(ii) In the second step, carbon monoxide is produced by the reaction of carbon dioxide with carbon.



Choose the correct statement about this reaction.

Tick (✓) **one** box.

the carbon dioxide is oxidised and the carbon is reduced

☐

both carbon dioxide and carbon are oxidised

☐

the carbon dioxide is reduced and the carbon is oxidised

☐

both carbon dioxide and carbon are reduced

☐

[1]

(iii) In the third step, iron(III) oxide is reduced by carbon monoxide.
The reaction is exothermic.

State the meaning of the term exothermic.

.....

..... [2]

(c) Calcium carbonate is added to the blast furnace.
The calcium carbonate breaks down as shown.



(i) Name the type of chemical reaction that takes place.

..... [1]

(ii) Complete this sentence about the calcium oxide that is produced in the blast furnace.

Calcium oxide reacts with impurities in the iron ore to form [1]

(d) Table 5.2 gives the observations when four different metals react with air.

Table 5.2

metal	observations
cerium	forms an oxide layer slowly without heating
copper	forms an oxide layer only when heated
gold	does not form an oxide layer even when heated
rubidium	forms an oxide layer quickly without heating

Put the four metals in order of their reactivity.

Put the least reactive metal first.

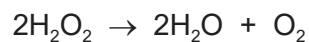
least reactive  most reactive

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[2]

[Total: 13]

- 6 Hydrogen peroxide, H_2O_2 , breaks down slowly at 40°C to produce oxygen gas and water.



A student investigates the breakdown of hydrogen peroxide at 40°C in the presence of a catalyst.

- (a) Fig. 6.1 shows the volume of oxygen gas released as the reaction proceeds.

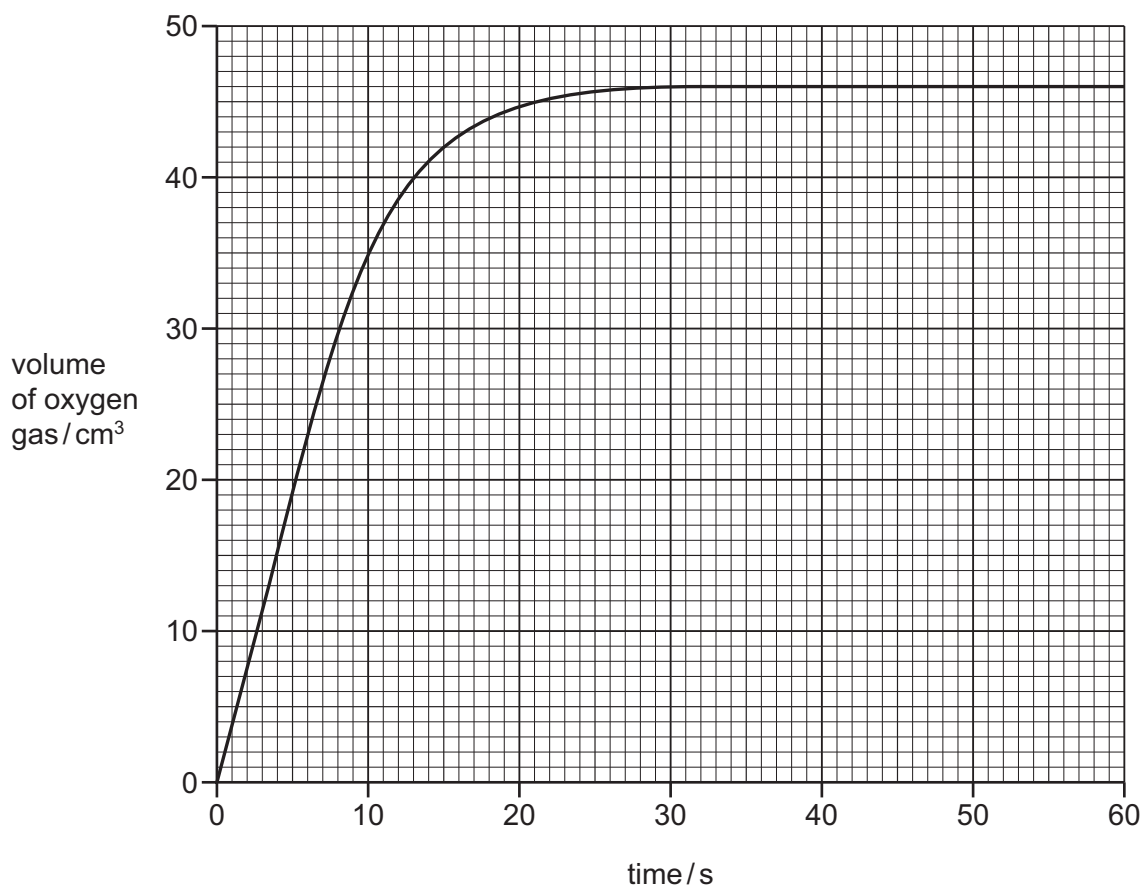


Fig. 6.1

- (i) Deduce the volume of oxygen gas released after 15 seconds.

volume of oxygen = cm^3 [1]

- (ii) The student repeats the experiment at 20°C .

All other conditions stay the same.

Draw a line on the grid in Fig. 6.1 to show how the volume of oxygen changes when a temperature of 20°C is used. [2]

- (b) (i)** The student repeats the experiment without a catalyst.

All other conditions stay the same.

Describe how the rate of reaction differs when no catalyst is used.

..... [1]

- (ii)** The student repeats the experiment using a lower concentration of hydrogen peroxide.

All other conditions stay the same.

Describe how the rate of reaction differs when a lower concentration of hydrogen peroxide is used.

..... [1]

- (c)** Hydrogen peroxide can act as a reducing agent in the presence of an alkali.

- (i)** State the meaning of the term alkali.

..... [1]

- (ii)** Give the formula of the ion that is present in all alkaline solutions.

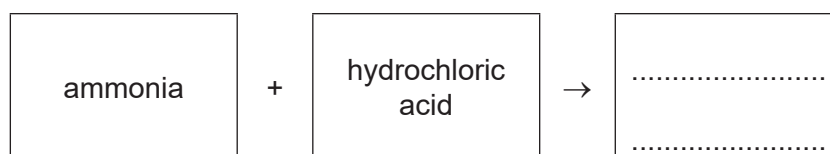
..... [1]

- (iii)** State the colour of methyl orange in an alkaline solution.

..... [1]

- (iv)** Aqueous ammonia is an alkali.

Complete the word equation for the reaction of aqueous ammonia with hydrochloric acid.

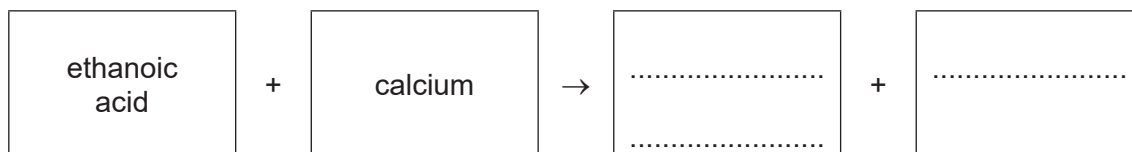


[1]

[Total: 9]

7 Ethanoic acid and methacrylic acid are both carboxylic acids.

(a) Complete the word equation for the reaction of ethanoic acid with calcium.



[2]

(b) Ethanoic acid can be reduced to ethanol.

(i) Name the homologous series that includes ethanol.

..... [1]

(ii) Ethanol can be manufactured by fermentation.

Describe **two** conditions needed for fermentation.

1

2

[2]

(c) Fig. 7.1 shows the displayed formula of methacrylic acid.

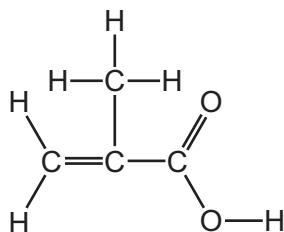


Fig. 7.1

(i) On Fig. 7.1, draw a circle around the functional group which reacts with aqueous bromine.

[1]

(ii) State the colour of aqueous bromine.

..... [1]

(iii) Deduce the molecular formula of methacrylic acid.

..... [1]

- (d) Methacrylic acid can be converted to methyl methacrylate.
The molecular formula of methyl methacrylate is $C_5H_8O_2$.

Complete Table 7.1 to calculate the relative molecular mass of methyl methacrylate.

Table 7.1

atom	number of atoms	relative atomic mass	
carbon	5	12	$5 \times 12 = 60$
hydrogen		1	
oxygen		16	

relative molecular mass = [2]

- (e) Methyl methacrylate can be polymerised to produce a plastic.

Describe **two** environmental problems caused by plastics.

1

2 [2]

- (f) Poly(ethene) is a polymer.

Draw the displayed formula of the monomer used to make poly(ethene).

[1]

[Total: 13]

8 Potassium chloride is an ionic compound.

(a) Complete Fig. 8.1 to show:

- the electronic configuration of a potassium ion
- the charge on the ion.

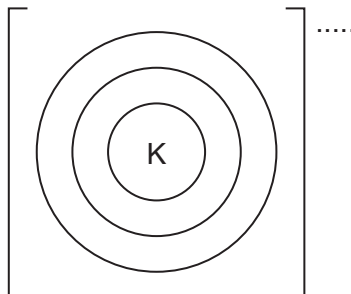


Fig. 8.1

[2]

(b) Deduce the number of protons and neutrons in the chloride ion shown.



number of protons

number of neutrons

[2]

(c) Molten potassium chloride is electrolysed using graphite electrodes.

(i) Define the term electrolysis.

.....
 [2]

(ii) State the names of the products at each electrode and give the observations at the positive electrode.

product at the negative electrode

product at the positive electrode

observations at the positive electrode

.....
 [3]

(d) Graphite electrodes are inert.

Name one **other** inert electrode.

..... [1]

(e) Graphite and diamond are two forms of carbon.

Fig. 8.2 shows the structure of diamond.

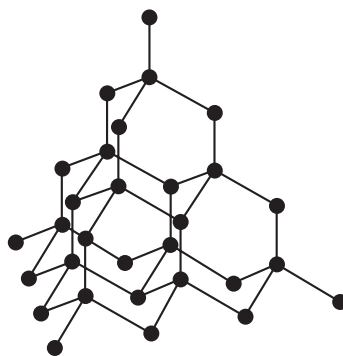


Fig. 8.2

(i) Name the type of bonding in diamond.

..... [1]

(ii) Use Fig. 8.2 to explain why diamond is used in cutting tools.

..... [1]

[Total: 12]

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Group																	
I	II							III	IV	V	VI	VII	VIII				
		Key atomic number atomic symbol name relative atomic mass							1Hhydrogen1								
3Li lithium 7	4Be beryllium 9									5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	10Ne neon 20		
11Na sodium 23	12Mg magnesium 24									13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	18Ar argon 40		
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131
55Cs caesium 133	56Ba barium 137	57–71lanthanoids	72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	86Rn radon —
87Fr francium —	88Ra radium —	89–103actinoids	104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —	113Nh nihonium —	114Fl flerovium —	115Mc moscovium —	116Lv livermorium —	117Ts tennessine —	118Og oganesson —

lanthanoids

actinoids

57	La lanthanum 139	58	Ce cerium 140	59	Pr praseodymium 141	60	Nd neodymium 144	61	Pm promethium —	62	Sm samarium 150	63	Eu europium 152	64	Gd gadolinium 157	65	Tb terbium 159	66	Dy dysprosium 163	67	Ho holmium 165	68	Er erbium 167	69	Tm thulium 169	70	Yb ytterbium 173	71	Lu lutetium 175
89	Ac actinium —	90	Th thorium 232	91	Pa protactinium 231	92	U uranium 238	93	Np neptunium —	94	Pu plutonium —	95	Am americium —	96	Cm curium —	97	Bk berkelium —	98	Cf californium —	99	Es einsteinium —	100	Fm fermium —	101	Md mendelevium —	102	No nobelium —	103	Lr lawrencium —