



CANDIDATE
NAME

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CENTRE
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CANDIDATE
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0653/42

February/March 2020

1 hour 15 minutes

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Blank pages are indicated.

- 1 (a) Fig. 1.1 shows an incomplete food web of some organisms in a field.

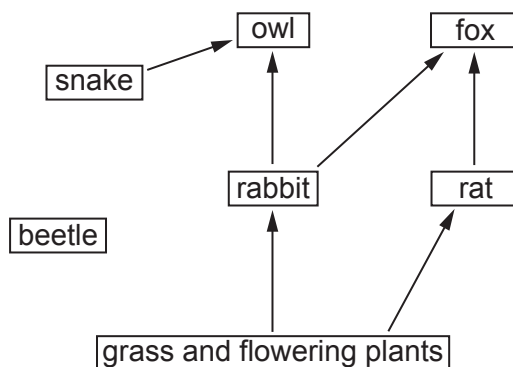


Fig. 1.1

- (i) The beetle feeds on grass and flowering plants.
The snake and owl both feed on the beetle.

Use this information to draw arrows to complete the food web in Fig. 1.1. [2]

- (ii) Name an organism in the food web in Fig. 1.1 which feeds at the same trophic level as the fox.

Explain your answer.

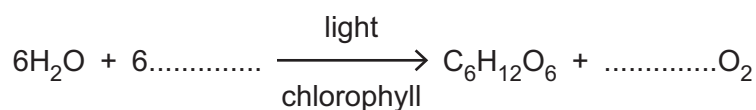
organism

explanation

..... [2]

- (b) The grass and flowering plants make their own organic nutrients during photosynthesis.

- (i) Complete the balanced symbol equation for photosynthesis.



[1]

- (ii) Describe the role of chlorophyll in the production of glucose during photosynthesis.

.....

.....

..... [2]

- (iii) Many glucose molecules may be joined together to form starch.

State the function of starch in plants.

..... [1]

- (iv) Name **one other** large molecule made from many glucose molecules joined together.

..... [1]

[Total: 9]

- 2 (a) The formula of methanol is CH_3OH .

Complete the structure of a molecule of methanol. Include all atoms and all bonds.



[1]

- (b) (i) Explain why methanol is **not** an alkane.

.....
 [1]

- (ii) Explain why a molecule of methanol contains **only** covalent bonds.

.....
 [1]

- (c) Methanol can be used as a fuel. It burns in oxygen and releases energy.

- (i) Complete the equation for the **complete** combustion of methanol.



[2]

- (ii) The combustion of methanol is an exothermic reaction.

Describe what is meant by *exothermic*.

Use ideas about energy, bond breaking and bond forming in your answer.

.....

 [2]

- (d) Methane is a greenhouse gas.

State **one** effect of increased levels of greenhouse gases in the atmosphere.

.....
 [1]

[Total: 8]

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- 3 Fig. 3.1 shows liquid (molten) iron being poured from a furnace into a mould to form an iron rod in the shape of a cylinder.

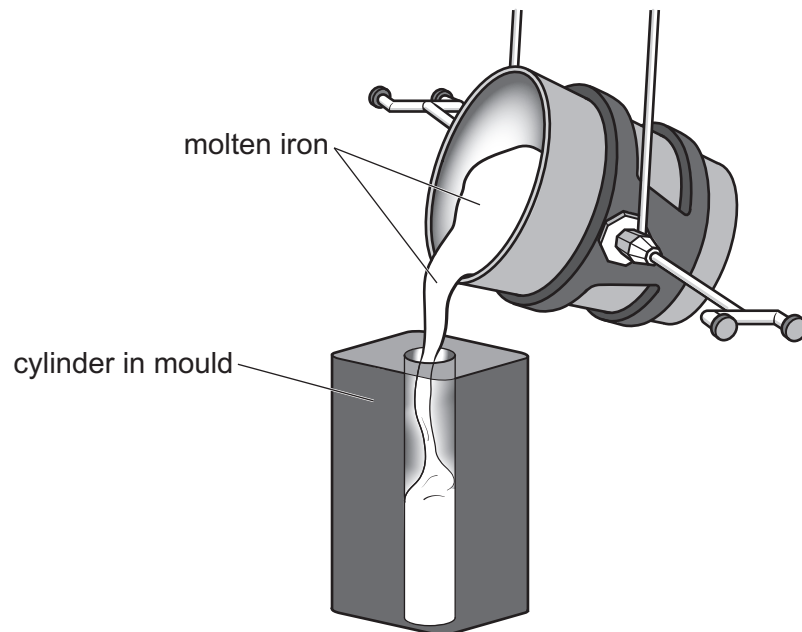


Fig. 3.1

Fig. 3.2 shows the iron rod after it has cooled down.



Fig. 3.2

- (a) (i) The bar is 80.0 cm long and 10.0 cm in diameter.

Show that the volume of the rod is 6280 cm³.

$$\pi = 3.14$$

[2]

- (ii) The density of solid iron is 7.86 g/cm^3 .

Calculate the mass of the rod in kilograms (kg).

mass = kg [3]

- (b) The density of molten iron is 6.98 g/cm^3 .

Explain why the density of molten iron is lower than the density of solid iron.

.....
.....
.....
.....
..... [2]

- (c) The workers near the furnace wear special clothing to protect them against the infrared radiation from the molten iron.

Explain why the clothing is coated with metal foil to give the best protection.

.....
.....
..... [2]

[Total: 9]

4 (a) Food taken in by the body must be digested so that the body can absorb the nutrients.

(i) Define the term *mechanical digestion*.

.....

 [2]

(ii) Describe **one** example of mechanical digestion in the human alimentary canal.

.....
 [1]

(b) Table 4.1 shows some information about chemical digestion in the human alimentary canal.

Table 4.1

digestive enzyme	function
lipase	breaks down into
protease	breaks down into

(i) State **one** part of the alimentary canal where **both** lipase **and** protease are secreted.

..... [1]

(ii) Complete Table 4.1.

[2]

(iii) State the part of the alimentary canal where food is ingested.

..... [1]

(c) Table 4.2 shows some statements about aerobic respiration.

In Table 4.2 place a tick (✓) to show the correct statements about aerobic respiration.

Table 4.2

statement about aerobic respiration	tick (✓) if correct
consists of chemical reactions	
energy released is used to relax muscles	
only takes place in animals	
produces oxygen	
produces water	
requires heat	

[2]

[Total: 9]

- 5 (a) A student places a spot of copper chloride solution onto a piece of chromatography paper, as shown in Fig. 5.1.

She places the chromatography paper in a solvent until the solvent front reaches the position shown in Fig. 5.2.

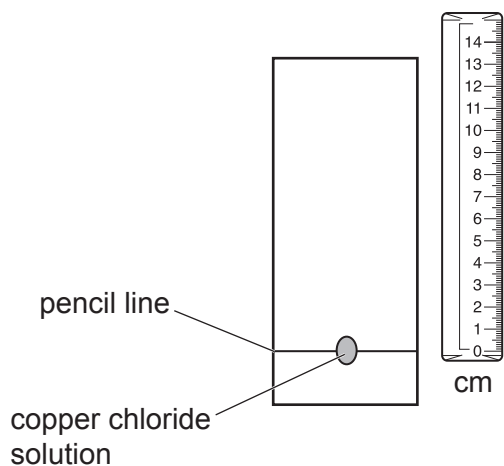


Fig. 5.1

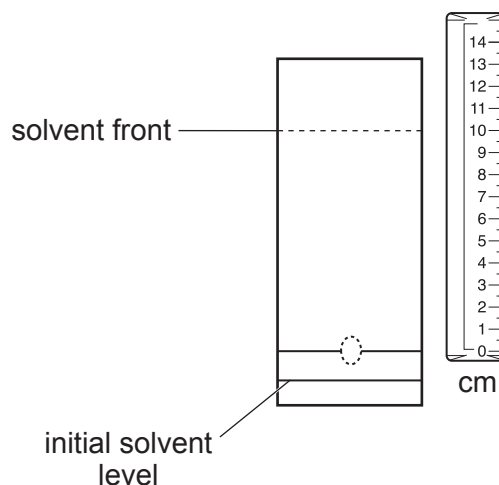


Fig. 5.2

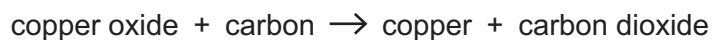
Copper chloride contains copper ions, Cu^{2+} . The R_f value for Cu^{2+} ions is 0.6.

On Fig. 5.2, draw a small circle to show the position of the Cu^{2+} ions.

[1]

- (b) Copper is extracted from copper oxide by heating with carbon.

- (i) The word equation for this reaction is shown.



State whether copper oxide is oxidised or reduced in this reaction.

Explain your answer.

.....

 [1]

- (ii) Explain why sodium oxide does **not** react with heated carbon.

.....
 [1]

(c) Copper is also extracted by electrolysis of molten copper chloride.

(i) State **one other** product obtained during this electrolysis.

..... [1]

(ii) Describe what happens to a copper ion, Cu^{2+} , at the cathode during this electrolysis.

Use ideas about electrons in your answer.

.....
.....
..... [2]

(d) Bronze is an alloy of copper and tin. Bronze is used to make coins.

Explain why pure copper is **not** used to make coins.

..... [1]

[Total: 7]

- 6 Fig. 6.1 shows a crane lifting a load to the top of a building. The crane uses an electric motor to lift the load.

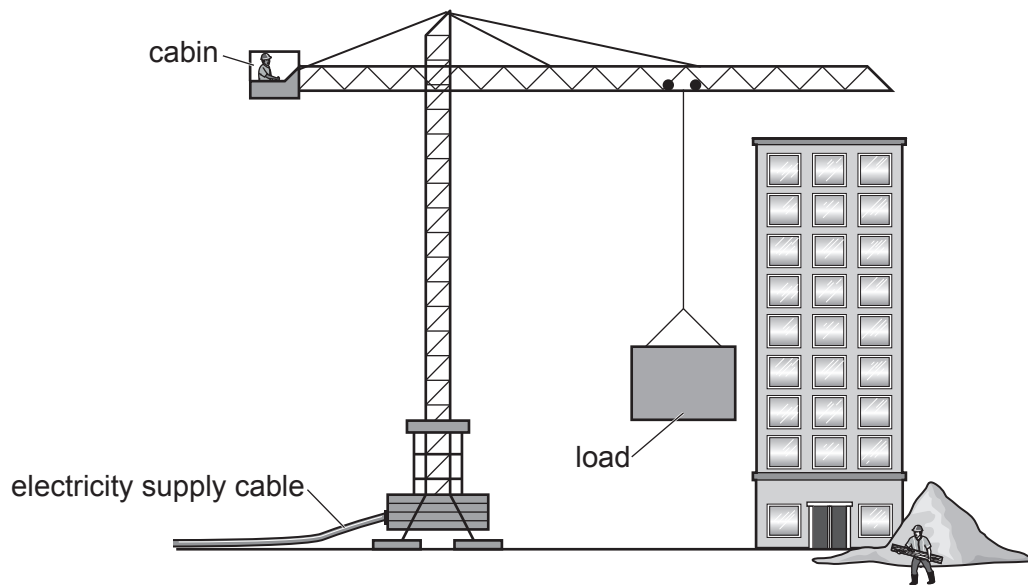


Fig. 6.1

- (a) At the start, the load is at rest on the ground. The load is lifted at a constant acceleration for 2.0 s.

At 2.0 s the load is moving upwards at a constant speed of 0.50 m/s.

Calculate the acceleration of the load during the first 2.0 s and give the unit.

acceleration = unit [3]

(b) The mass of the load is 500 kg.

- (i) The load is lifted from the ground to the top of the building 25 m above the ground.
Gravitational field strength is 10 N/kg.

Calculate the work done on the load.

Show your working.

work done = J [2]

- (ii) The power of the electric motor lifting the load is 5 kW.
The crane takes 56 s to lift the load to the top of the building.

Calculate the electrical energy supplied to the electric motor in this time.

energy = J [2]

- (iii) The electrical energy supplied is greater than the useful work done on the load. Some electrical energy is transferred in other ways.

Suggest **one** other way in which this electrical energy is used.

.....

 [1]

[Total: 8]

- 7 (a) Potato pieces **A**, **B** and **C** are immersed in salt solutions of different concentrations for 30 minutes.

The lengths of the potato pieces are measured before and after immersion.

Table 7.1 shows the results.

Table 7.1

potato piece	concentration of salt solution/ mol per dm ³	length of potato piece before immersion/mm	length of potato piece after immersion/mm
A	0.1	80	85
B	0.4	80	80
C	0.6	80	76

- (i) Calculate the percentage increase in length of potato piece **A**.

increase =% [2]

- (ii) Explain why the length of potato piece **C** decreases.

Use the term *osmosis* in your answer.

.....

 [3]

- (b) Suggest why it is an advantage for a plant to have longer root hairs in areas where the soil is dry.

.....

 [2]

(c) Mineral ions are essential for the healthy growth of plants.

(i) State the function of magnesium ions in plants.

.....
..... [1]

(ii) Describe the effect of a deficiency in magnesium ions in plants.

..... [1]

[Total: 9]

- 8 (a) A bromine atom contains seven electrons in its outer shell.

Complete the dot-and-cross diagram of a molecule of bromine in Fig. 8.1. Show only the outer shell electrons.



Fig. 8.1

[2]

- (b) The boiling point of liquid bromine is 59 °C.

Explain what happens to bromine molecules when liquid bromine boils.

Use ideas about molecules and energy changes in your answer.

.....

.....

.....

..... [2]

- (c) Orange aqueous bromine is mixed with colourless aqueous potassium chloride.

Predict the colour change, if any, in this reaction.

Explain your answer.

colour change

explanation

..... [2]

(d) When large hydrocarbon molecules are cracked, alkenes are produced.

(i) State **two** conditions needed for this process.

1.

2. [2]

(ii) Aqueous bromine is used to test for the presence of alkenes.

State the colour change that indicates a positive result for this test.

from to [1]

(e) The formula of aluminium bromide is Al_2Br_6 .

The melting point of aluminium bromide is 97°C and the boiling point is 255°C .

Explain how this information suggests that aluminium bromide contains covalent bonds.

Use ideas about attractive forces in your answer.

.....

.....

..... [2]

[Total: 11]

- 9 Fig. 9.1 shows an ambulance. On the roof it has a flashing blue lamp and a siren.



Fig. 9.1

- (a) A car driver sees the ambulance appear in the far distance.
The driver hears the siren 3 s after seeing the blue light.

Explain why the driver sees the blue light before he hears the siren.

.....
..... [1]

- (b) The siren emits sounds at two frequencies, 600 Hz and 1500 Hz.
The siren has a plastic casing. The sound travels through the plastic casing at a speed of 2200 m/s.

Calculate the wavelength of the 600 Hz sound.

wavelength = m [2]

- (c) An electric motor rotates a mirror around the blue lamp to reflect a bright beam of light.

The motor has a power rating of 20 W and the lamp has a power rating of 60 W.

The lamp and the motor both operate at a potential difference (p.d.) of 12 V from the 12 V ambulance battery.

- (i) State the type of circuit arrangement required for the motor and the lamp.

..... [1]

- (ii) Calculate the current in the motor when turning the mirror.

current = A [2]

- (iii) The lamp and motor are switched on by one switch even if the siren is not being used.

The siren has its own switch and also operates at 12 V.

On Fig. 9.2 complete the diagram for the circuit that operates the lamp, the motor and the siren from the 12 V car battery.

The circuit symbol for a siren is 

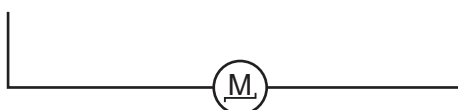
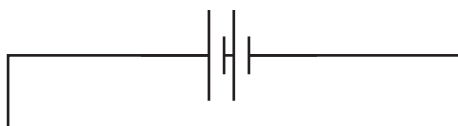


Fig. 9.2

[4]

[Total: 10]

The Periodic Table of Elements

[illegible]

lanthanoids	57	La	lanthanum	139	58	Ce	cerium	140	59	Pr	praseodymium	141	60	Nd	neodymium	144	61	Pm	promethium	—	62	Sm	samarium	150	63	Eu	europlum	152	64	Gd	gadolinium	157	65	Tb	terbium	159	66	Dy	dysprosium	163	67	Ho	holmium	165	68	Er	erbium	167	69	Tm	thulium	169	70	Yb	ytterbium	173	71	Lu	lutetium	175	
	actinoids	89	Ac	actinium	—	90	Th	thorium	232	91	Pa	protactinium	231	92	U	uranium	238	93	Np	neptunium	—	94	Pu	plutonium	—	95	Am	americium	—	96	Cm	curium	—	97	Bk	berkelium	—	98	Cf	californium	—	99	Es	einsteinium	—	100	Fm	fermium	—	101	Md	mendelevium	—	102	No	nobelium	—	103	Lr	lawrencium	—

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).