



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0653/31

October/November 2020

1 hour 15 minutes

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Blank pages are indicated.

- 1 (a) Fig. 1.1 is a diagram of some parts of the body.

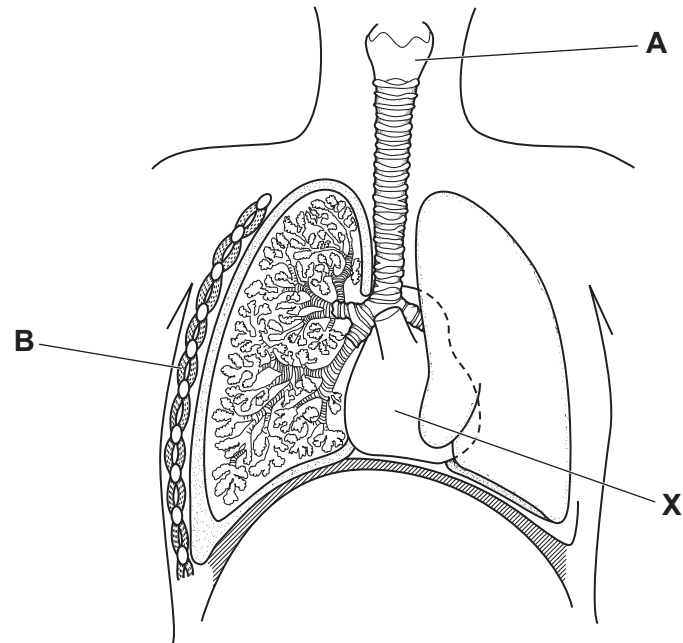


Fig. 1.1

- (i) Identify parts **A** and **B** shown in Fig. 1.1.

A

B [2]

- (ii) Describe the function of part **X** shown in Fig. 1.1.

.....

..... [1]

(b) Fig. 1.2 shows the rate and depth of breathing of a person while at rest and during exercise.

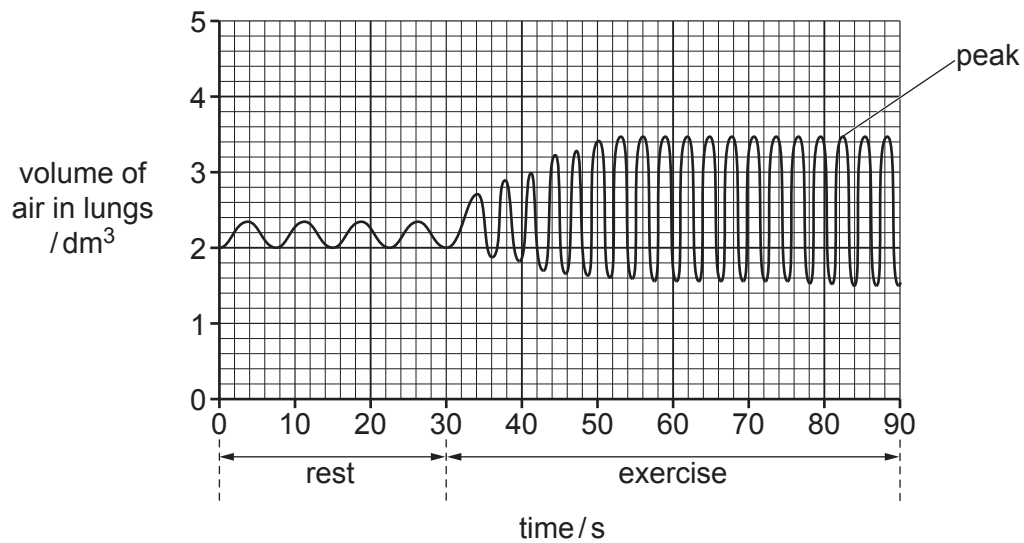


Fig. 1.2

Each time the line rises to a peak, the person is breathing in.

(i) Use Fig. 1.2 to determine the number of breaths the person takes while at rest.

number of breaths = [1]

(ii) Describe how the rate and depth of breathing of the person change during exercise, as shown in Fig. 1.2.

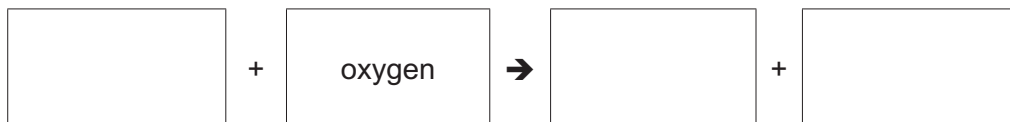
rate

depth

[2]

(c) The purpose of breathing is to obtain oxygen for respiration.

Complete the word equation for respiration.



[2]

[Total: 8]

- 2 (a) Copper is extracted from copper oxide by heating with carbon, as shown in Fig. 2.1.

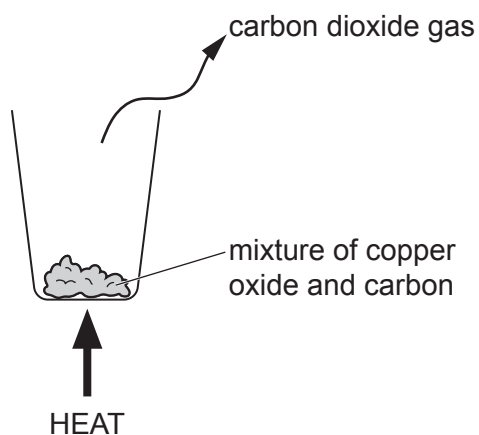
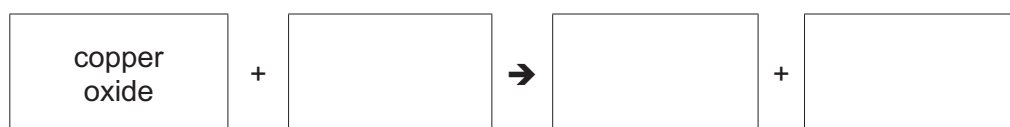


Fig. 2.1

Carbon dioxide is produced in this process.

- (i) Complete the word equation for this process.



[2]

- (ii) Use words from the list to describe substances involved in this process.

Each word may be used once, more than once or not at all.

atoms

compound

covalent

element

ionic

ions

solution

Copper oxide is a containing oppositely charged

In carbon dioxide, each molecule contains three
chemically joined by bonds.

[4]

- (iii) State whether the change from copper oxide to copper during this process is oxidation or reduction.

Explain your answer.

change

explanation

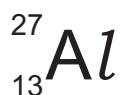
..... [1]

- (b) Aluminium is extracted from the ore bauxite.

- (i) State the method of extraction of aluminium from bauxite.

..... [1]

- (ii) An atom of aluminium is represented by the symbol shown.



Deduce the number of electrons and the number of neutrons in this atom.

electrons

neutrons

[2]

[Total: 10]

- 3 (a) Fig. 3.1 shows the distance–time graph for a man.

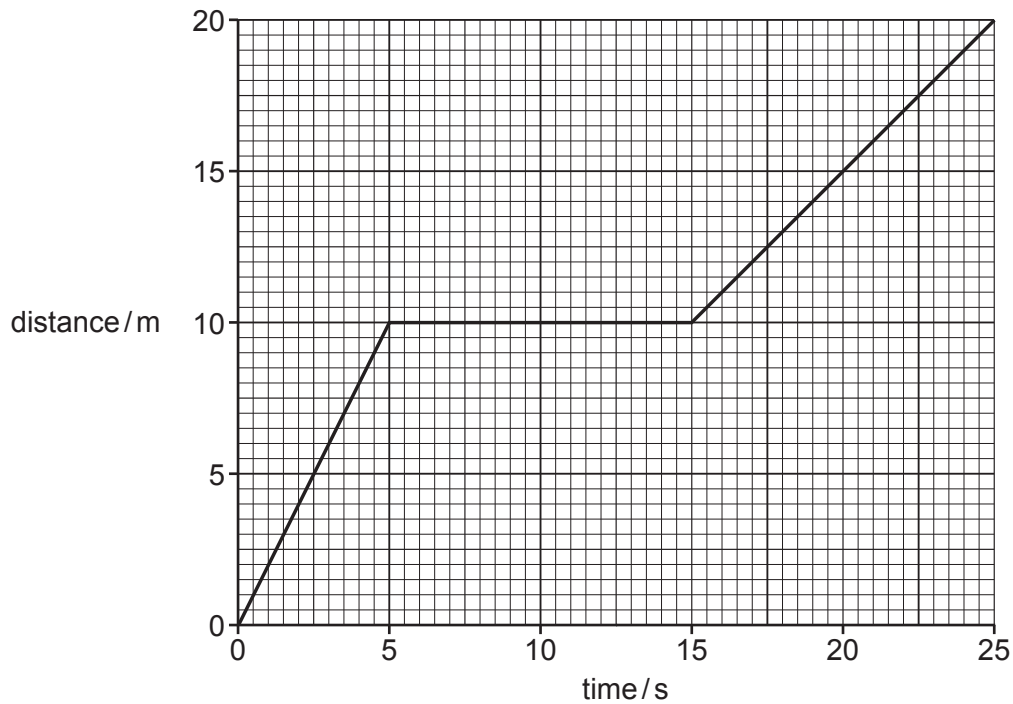


Fig. 3.1

- (i) Suggest what the man is doing between 5 s and 15 s.
 [1]
- (ii) On Fig. 3.1, draw an **X** on the graph to show when the man is moving fastest. [1]
- (iii) Use Fig. 3.1 to calculate the average speed of the man for the 25 s.

speed = m/s [2]

- (b) The weight of the man is 800 N.

The gravitational field strength g is 10 N/kg.

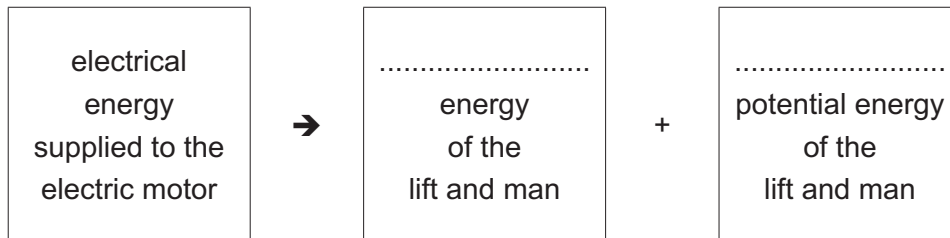
Calculate the mass of the man.

mass = kg [1]

- (c) (i) The man enters a lift (elevator). The lift moves the man vertically upwards.

The lift uses an electric motor.

Complete the useful energy transfers for the lift and man.



[2]

- (ii) The amount of electrical energy supplied to the electric motor is actually greater than the useful work done in moving the lift and man up to the higher level.

Suggest why.

.....

.....

..... [2]

[Total: 9]

- 4 (a) Fig. 4.1 shows diagrams of a plant cell before and after the cell is immersed in concentrated sugar solution.

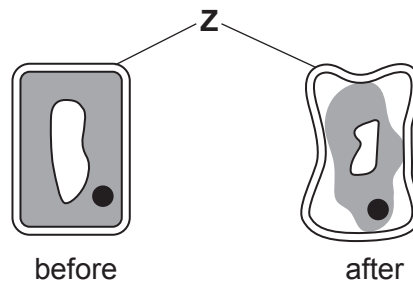


Fig. 4.1

- (i) Identify part **Z** shown in Fig. 4.1.

..... [1]

- (ii) Describe the effect of immersing the cell in concentrated sugar solution.

.....
.....
.....
..... [2]

(b) Fig. 4.2 is a drawing of a cross-section of a plant root.

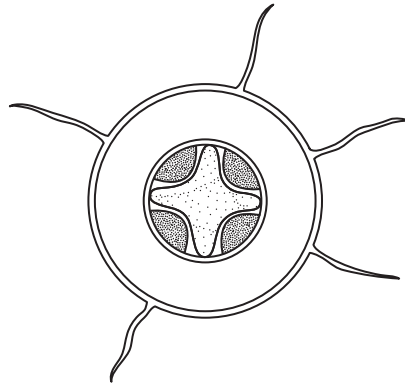


Fig. 4.2

(i) On Fig. 4.2, use a label line and the letter **X** to show the position of the xylem. [1]

(ii) State the function of root hair cells.

.....
 [1]

(iii) Water moves through xylem vessels to the leaves due to transpiration.

State the process by which water in the leaves is lost:

from the surface of the mesophyll cells
 through the stomata. [2]

(c) Water is used in photosynthesis.

Photosynthesis is the process by which plants manufacture carbohydrates from raw materials using energy from light.

(i) State the name of the structures inside plant cells where photosynthesis takes place.

..... [1]

(ii) State the name of the gas produced in photosynthesis.

..... [1]

[Total: 9]

5 (a) Air is a mixture of different gases.

(i) Name the gas that makes up 78% of clean air.

..... [1]

(ii) Name the gas that must be present for iron to rust.

..... [1]

(iii) Name **one** common air pollutant and describe **one** adverse effect that is caused by this pollutant.

pollutant

adverse effect [2]

(iv) State the formulae of **two** greenhouse gases.

1

2 [2]

(b) Fig. 5.1 is an incomplete dot-and-cross diagram of a molecule of water.

On Fig. 5.1, draw dots and crosses to show **all** of the outer shell electrons in a molecule of water.

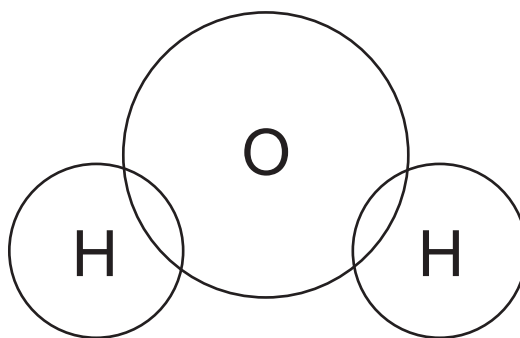


Fig. 5.1

[3]

(c) Describe the effect of water on blue cobalt(II) chloride.

..... [1]

[Total: 10]

- 6 Fig. 6.1 shows a girl using a bicycle pump to add air to a bicycle tyre.

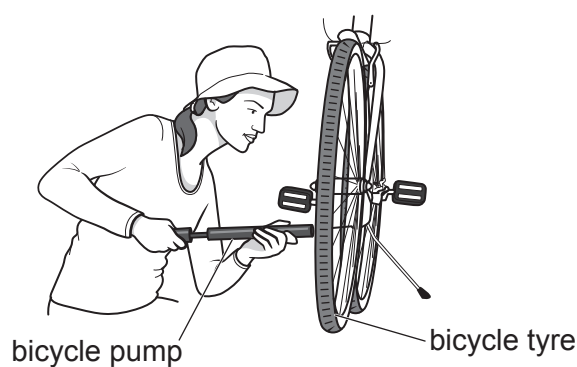


Fig. 6.1

- (a) Fig. 6.2 shows the arrangement of molecules in the air inside the bicycle tyre.

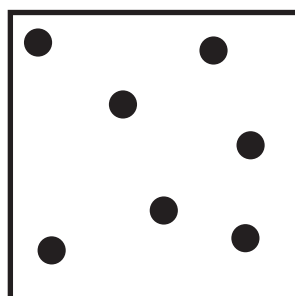


Fig. 6.2

- (i) Suggest what happens to the separation of the molecules as the girl pumps more air into the tyre.

.....
 [1]

- (ii) The temperature of the air in the tyre increases.

Describe how the movement of the molecules changes as this happens.

.....
 [1]

(b) Fig. 6.3 shows the girl wearing a cotton hat on a sunny day.

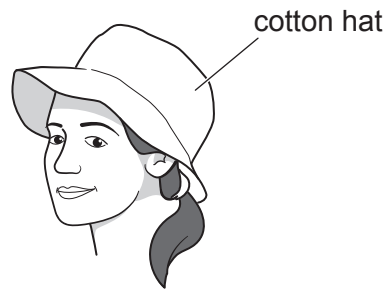


Fig. 6.3

(i) State the method of energy transfer from the Sun to the Earth.

..... [1]

(ii) Suggest **two** ways the cotton hat reduces energy transfer from the Sun to the girl's head.

1

.....

2

.....

[2]

(iii) The girl puts a sunscreen cream on the skin of her face.

Suggest how this precaution helps reduce the risk of sunburn.

.....

.....

..... [2]

(c) Fig. 6.4 shows a bell fixed on the handlebars of the bicycle.

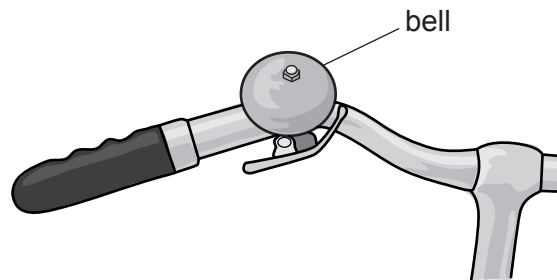


Fig. 6.4

(i) Fig. 6.5 represents the sound wave produced when the girl rings the bell.

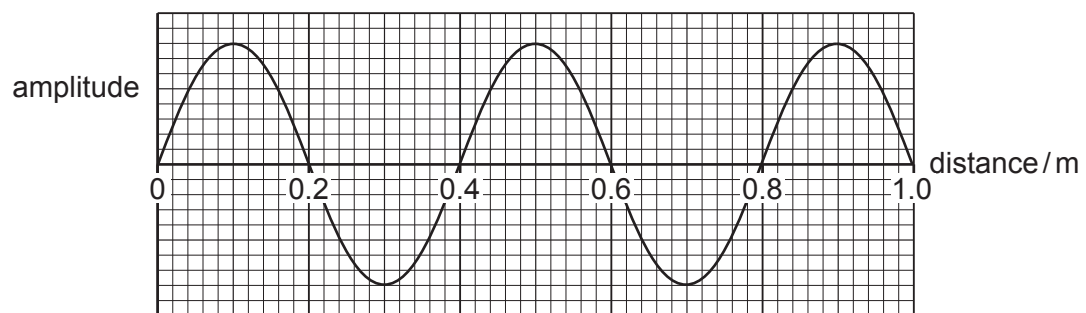


Fig. 6.5

Use Fig. 6.5 to find the wavelength of the sound wave produced by the bell.

wavelength = m [1]

(ii) The girl makes the bell ring louder.

Describe the difference this makes to the sound wave produced.

..... [1]

[Total: 9]

- 7 (a) The boxes on the left show some parts of the male reproductive system in humans. The boxes on the right show the function of each part.

Draw **one** straight line from each part to its function.

part	function
penis	production of sperm
prostate gland	secretion of fluid for sperm to swim in
testes	transfer of semen to vagina during sexual intercourse

[2]

- (b) Fig. 7.1 is a diagram of the female reproductive system in humans.

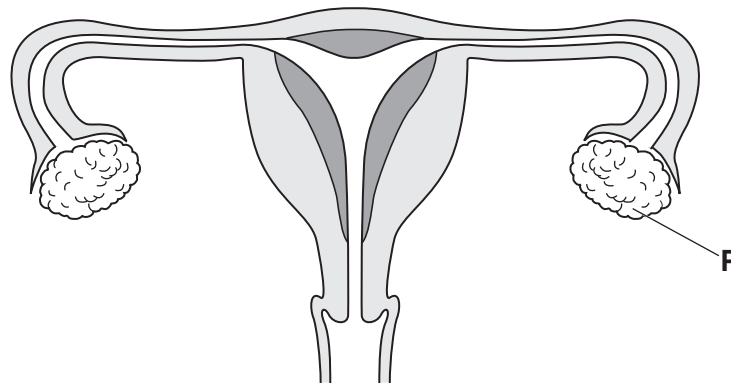


Fig. 7.1

- (i) On Fig. 7.1, use a label line and the letter **F** to show where fertilisation occurs.

[1]

(ii) Describe the changes that occur in structure **P** during the menstrual cycle.

.....

.....

.....

..... [2]

(iii) Fig. 7.2 shows the process of human fertilisation and early development of an embryo.

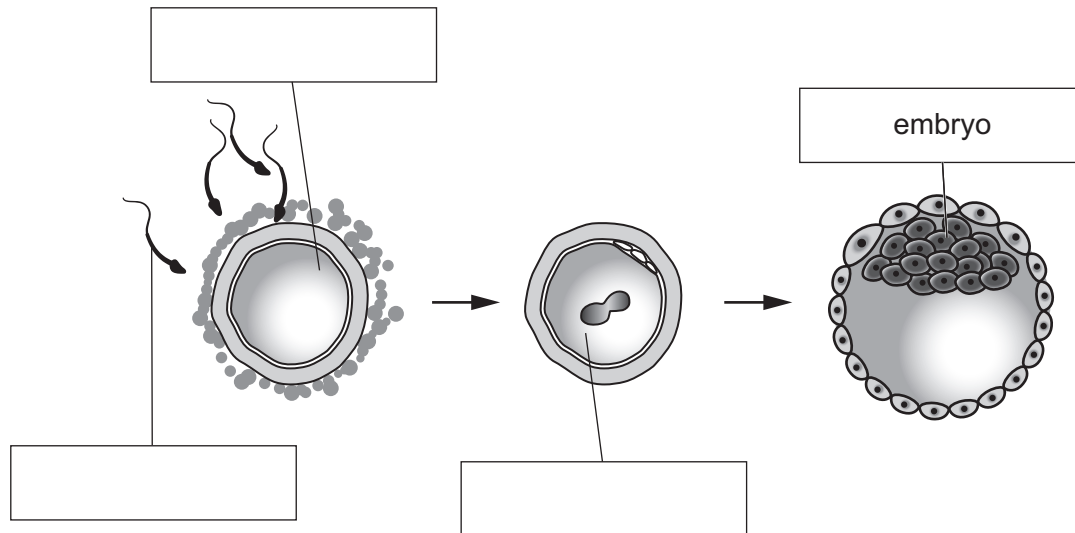


Fig. 7.2

Complete Fig. 7.2 using words from the list.

fetus	nucleus	ovum
sperm	uterus	zygote

[3]

(c) Human immunodeficiency virus (HIV) is a sexually transmitted infection.

State **two** ways of preventing the transmission of HIV.

1

2 [2]

[Total: 10]

- 8 (a) Sodium chloride is made when aqueous sodium hydroxide is mixed with dilute hydrochloric acid.

State what happens to the pH of the mixture as the aqueous sodium hydroxide is added to the dilute hydrochloric acid.

..... [1]

- (b) Sodium is in Group I of the Periodic Table, and chlorine is in Group VII.

- (i) Describe the change in character of elements across a period in the Periodic Table from left to right.

.....

..... [1]

- (ii) Describe the trend in reactivity of Group I metals from lithium to potassium.

.....

..... [1]

- (iii) Describe the trend in physical state of Group VII elements from fluorine to iodine.

.....

..... [1]

- (c) Aqueous bromine reacts with alkenes.

- (i) State the colour change that occurs during this reaction.

..... [1]

- (ii) Name the reaction that produces alkenes from larger alkane molecules.

..... [1]

[Total: 6]

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- 9 Fig. 9.1 shows an electric toaster that is used for toasting slices of bread.

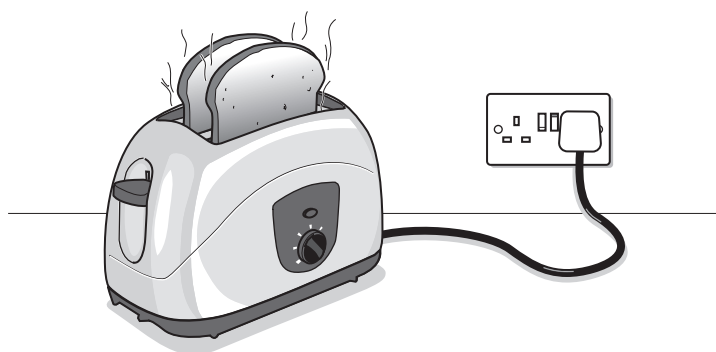


Fig. 9.1

- (a) The toaster uses a 230 V mains electricity supply as a source of energy.

The toaster contains:

- two heaters connected in parallel
- one switch to turn both heaters on and off
- a fuse to protect the circuit.

On Fig. 9.2, complete the circuit diagram for the toaster.

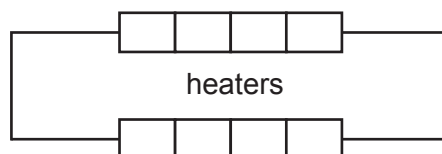
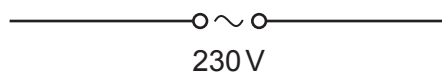


Fig. 9.2

[4]

(b) A wire in one of the heaters breaks, and that heater stops working. The fuse is not affected.

(i) Explain why the other heater continues to work.

.....
 [1]

(ii) The potential difference (p.d.) across the working heater is 230 V.

The current in the working heater is 2.5 A.

Calculate the resistance of the heater.

Give the unit of your answer.

resistance = unit [3]

(c) The fuse in the circuit now needs replacing. The current from the source when both heaters are working is 5.0 A.

Circle the most appropriate fuse rating for the replacement fuse. Give a reason for your choice.

1 A 5 A 13 A

reason
 [1]

[Total: 9]

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Group																	
I	II	1 H hydrogen 1					III	IV	V	VI	VII	VIII					
		Key atomic number atomic symbol name relative atomic mass															
3 Li lithium 7	4 Be beryllium 9																
11 Na sodium 23	12 Mg magnesium 24																
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	114 Fl flerovium —	116 Lv livermorium —	117 Ts tennessine —	118 Og oganesson —	119 Uue unbinilium —	120 Uuh ununilium —

lanthanoids	57	La	lanthanum	139	58	Ce	cerium	140	59	Pr	praseodymium	141	60	Nd	neodymium	144	61	Pm	promethium	—	62	Sm	samarium	150	63	Eu	europium	152	64	Gd	gadolinium	157	65	Tb	terbium	159	66	Dy	dysprosium	163	67	Ho	holmium	165	68	Er	erbium	167	69	Tm	thulium	169	70	Yb	ytterbium	173	71	Lu	lutetium	175	
	actinoids	89	Ac	actinium	—	90	Th	thorium	232	91	Pa	protactinium	231	92	U	uranium	238	93	Np	neptunium	—	94	Pu	plutonium	—	95	Am	americium	—	96	Cm	curium	—	97	Bk	berkelium	—	98	Cf	californium	—	99	Es	einsteinium	—	100	Fm	fermium	—	101	Md	mendelevium	—	102	No	nobelium	—	103	Lr	lawrencium	—