



Cambridge IGCSE™

COMBINED SCIENCE

0653/13

Paper 1 Multiple Choice (Core)

October/November 2023

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **12** pages.



1 What is osmosis?

- A the movement of salt across a cell wall
- B the movement of salt across a partially permeable membrane
- C the movement of water across a cell wall
- D the movement of water across a partially permeable membrane

2 Which row shows the elements that make up proteins?

	carbon	hydrogen	nitrogen	oxygen
A	✓	✗	✓	✗
B	✓	✓	✗	✗
C	✗	✓	✗	✓
D	✓	✓	✓	✓

3 What are enzymes made from?

- A fat
- B protein
- C starch
- D oil

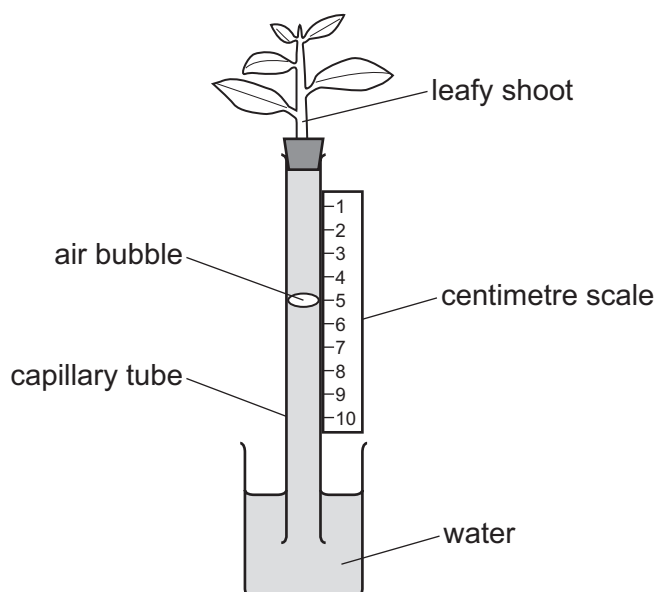
4 Which row correctly matches an organ of the alimentary canal with its functions?

	organ	absorption	digestion	egestion	ingestion
A	large intestine	yes	no	no	yes
B	oesophagus	no	yes	yes	no
C	small intestine	yes	yes	no	no
D	stomach	no	no	yes	yes

5 Which process is defined as taking substances into the body through the mouth?

- A absorption
- B digestion
- C egestion
- D ingestion

- 6 The apparatus shown can be used to measure the rate of transpiration.



Four identical sets of apparatus were set up under different environmental conditions and left for 1 hour.

In all four apparatus, the air bubble was at the 5.0 cm point at the start of the experiment.

What would be the reading on the scale of the apparatus that was left in low humidity and high temperature?

- A** 1.5 cm **B** 5.0 cm **C** 7.0 cm **D** 9.5 cm

- 7 A student is investigating the differences in composition of samples of inspired and expired air.

What can he use to test for carbon dioxide?

- A** biuret solution
B limewater
C ethanol
D iodine solution

- 8 What is the word equation for aerobic respiration?

- A** carbon dioxide + water → glucose + oxygen
B glucose → carbon dioxide + water
C oxygen + carbon dioxide → glucose + water
D glucose + oxygen → carbon dioxide + water

9 Which statement about the growth response of plant roots is correct?

- A They grow away from gravity and away from light.
- B They grow away from gravity and towards light.
- C They grow towards gravity and away from light.
- D They grow towards gravity and towards light.

10 Which row describes asexual reproduction?

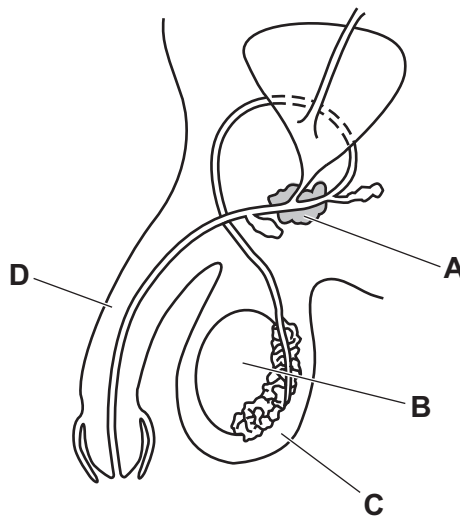
	number of parents	a zygote is produced	offspring genetically identical to the parent
A	1	no	yes
B	1	yes	no
C	2	no	yes
D	2	yes	no

11 Which statement describes fertilisation in a flowering plant?

- A fusion of a pollen nucleus with a nucleus in the ovule
- B fusion of a pollen nucleus with the stigma
- C transfer of a pollen grain from the anther to the stigma
- D transfer of a pollen grain from the filament to the stigma

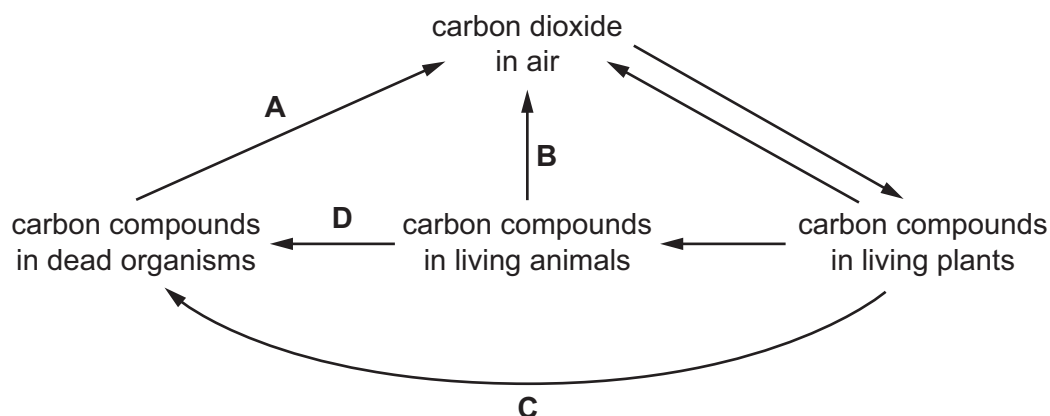
12 The diagram shows the human male reproductive system.

Which labelled part is the prostate gland?



13 The diagram shows part of the carbon cycle.

Which arrow represents respiration by decomposers?

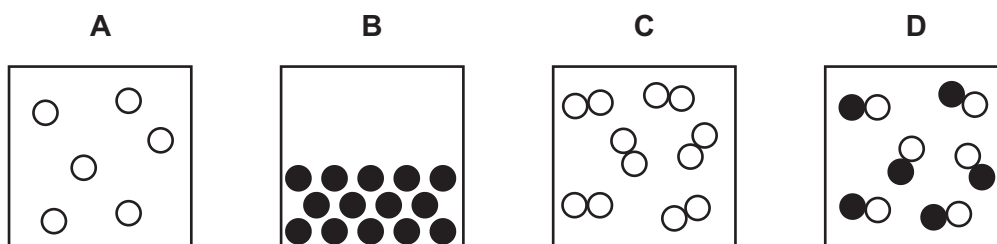


14 Substance X is an element.

It is a gas at room temperature.

It is made of X_2 molecules.

Which diagram represents X?



15 Which statement about tap water is correct?

- A** It is a compound.
- B** It is a mixture of elements.
- C** It is a pure substance.
- D** It is a solution.

16 Compound X contains one iron atom.

It also contains the same number of sulfur atoms as iron atoms and four times as many oxygen atoms as sulfur atoms.

What is the formula of compound X?

- A** $\text{Fe}(\text{SO})_4$
- B** FeSO_4
- C** FeS_4O
- D** $\text{Fe}_4\text{S}_4\text{O}$

17 When solid sodium carbonate and dilute hydrochloric acid are mixed, a reaction occurs.

During this reaction, carbon dioxide is released and the temperature of the mixture increases.

Which chemical terms describe this reaction?

- 1 exothermic
- 2 neutralisation
- 3 thermal decomposition

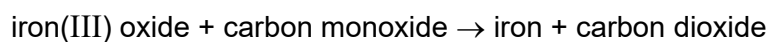
A 1, 2 and 3 **B** 1 and 2 only **C** 1 and 3 only **D** 2 and 3 only

18 In which row do **all** of the changes shown increase the rate of reaction?

	temperature	concentration	particle size
A	decrease	increase	decrease
B	decrease	decrease	increase
C	increase	increase	decrease
D	increase	decrease	increase

19 Iron is extracted from its ore using carbon monoxide.

The word equation is shown.

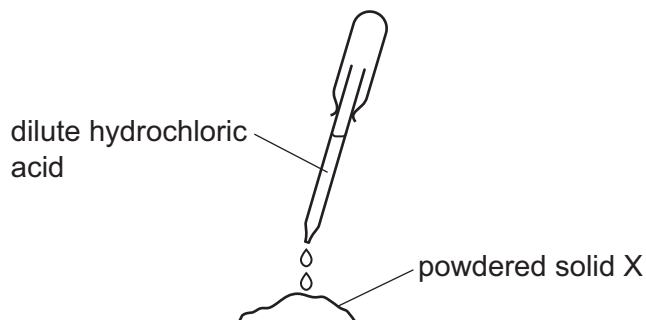


Which statement is correct?

- A** Carbon monoxide is oxidised by gaining oxygen.
- B** Carbon monoxide is reduced by losing oxygen.
- C** Iron(III) oxide is oxidised by losing oxygen.
- D** Iron(III) oxide is reduced by gaining oxygen.

- 20 Dilute hydrochloric acid is added to powdered solid X.

Hydrogen gas is produced.



What is X?

- A zinc
 - B zinc carbonate
 - C zinc hydroxide
 - D zinc oxide
- 21 Which test is used to identify ammonia?
- A A glowing splint relights.
 - B Damp blue litmus paper is bleached.
 - C Damp red litmus paper turns blue.
 - D Limewater turns milky.
- 22 Which statement about elements in Period 3 of the Periodic Table is correct?
- A All the elements in Period 3 are metals.
 - B All the elements in Period 3 are non-metals.
 - C Metals are on the left, non-metals are on the right.
 - D Non-metals are on the left, metals are on the right.

- 23** Hydrogen peroxide decomposes slowly to produce water and oxygen.

Element X and a compound of element X cause hydrogen peroxide to decompose more rapidly.

What is the position of element X in the Periodic Table?

A simplified periodic table grid is shown, consisting of 10 columns and 6 rows. The grid is shaped like a standard periodic table, with the first two columns on the left, a gap in the third column, and the remaining columns on the right. The letters A, B, C, and D are placed in specific cells: A is in the first column, second row; B is in the eighth column, second row; C is in the fifth column, third row; and D is in the tenth column, third row. An empty box is located above the gap in the third column, centered between the first and eighth columns.

- 24** Which substance does **not** react with dilute hydrochloric acid to produce copper chloride?

- A** copper
B copper carbonate
C copper hydroxide
D copper oxide

- 25** Which gases are greenhouse gases?

- A** carbon dioxide and methane
B carbon dioxide and oxygen
C methane and nitrogen
D nitrogen and oxygen

- 26** Decane is an alkane.

Which statement about decane is correct?

- A** It burns in air to form carbon dioxide and hydrogen.
- B** It is an unsaturated hydrocarbon.
- C** It only contains single C–C and C–H bonds.
- D** It rapidly decolourises bromine water.

27 Which process is used to produce alkenes?

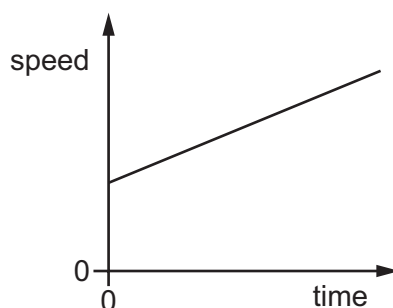
- A addition polymerisation
- B combustion
- C cracking
- D fractional distillation

28 An object travels 6.0 km in two minutes.

What is its speed?

- A 0.050 m/s B 3.0 m/s C 50 m/s D 3000 m/s

29 The graph shows how the speed of a moving car varies with time.



Which statement about the car is correct?

- A The car is accelerating.
- B The car is at rest at time = 0.
- C The car must be travelling in a straight line.
- D The car travels equal distances in equal times.

30 The gravitational field strength on the planet Mercury is 3.7 N/kg.

What is the weight of a 10 kg rock on Mercury?

- A 0.37 N B 3.7 N C 10 N D 37 N

31 A car has an initial kinetic energy of 120 kJ at the bottom of a slope. The car is driven up the slope. At the top of the slope, the car has 260 kJ of kinetic energy and has gained 570 kJ of gravitational potential energy.

What is the total increase in kinetic energy and gravitational potential energy of the car as it moves up the slope?

- A 430 kJ B 710 kJ C 830 kJ D 950 kJ

32 Which statement about the boiling point of a substance is correct?

- A** At all temperatures above its boiling point, a substance must be a gas.
- B** At all temperatures above its boiling point, a substance must be a liquid.
- C** At all temperatures below its boiling point, a substance must be a gas.
- D** At all temperatures below its boiling point, a substance must be a liquid.

33 A student puts an object made of metal and another object made of plastic in the same freezer for several days.

The student removes the two objects from the freezer.

When the student touches the objects, the metal one feels colder than the plastic one.

Why is this?

- A** The metal conducts heat quickly away from the hand.
- B** The metal is at a lower temperature than the plastic.
- C** The plastic conducts heat quickly into the hand.
- D** The plastic has a lower melting point than the metal.

34 What is a use of microwaves?

- A** checking for broken bones
- B** satellite television
- C** tanning lamps
- D** television remote controllers

35 A loudspeaker vibrates at different frequencies.

Which frequency of vibration does **not** produce a sound that a human can hear?

- A** 60 Hz **B** 600 Hz **C** 6.0 kHz **D** 60 kHz

36 A potential difference (p.d.) of 6.0 V is applied across a lamp.

The current in the lamp is 1.5 A.

What is the resistance of the lamp?

- A** 0.25 Ω **B** 4.0 Ω **C** 4.5 Ω **D** 9.0 Ω

- 37** A plastic rod is rubbed with a woollen cloth. The rod becomes negatively charged.

What happens to the woollen cloth?

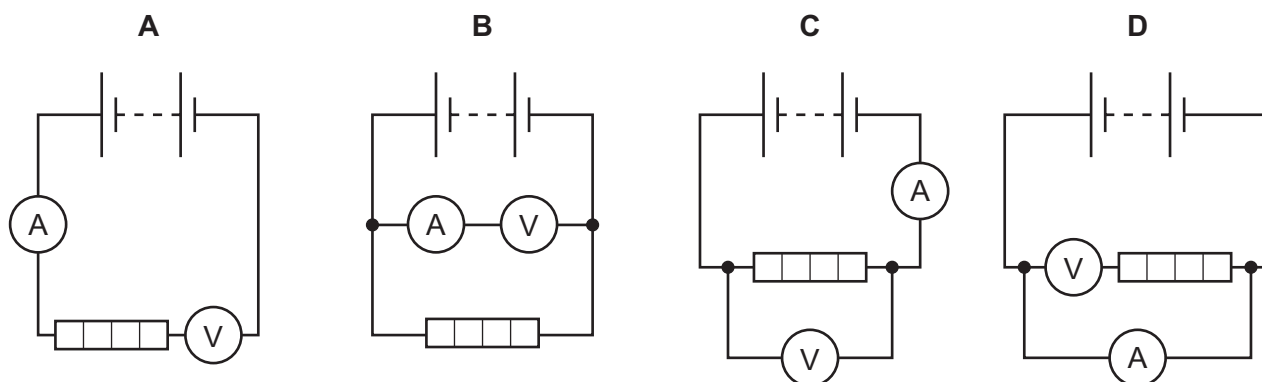
- A** It gains electrons and becomes negatively charged.
- B** It gains electrons and becomes positively charged.
- C** It loses electrons and becomes negatively charged.
- D** It loses electrons and becomes positively charged.

- 38** A battery is connected to a heater, an ammeter and a voltmeter.

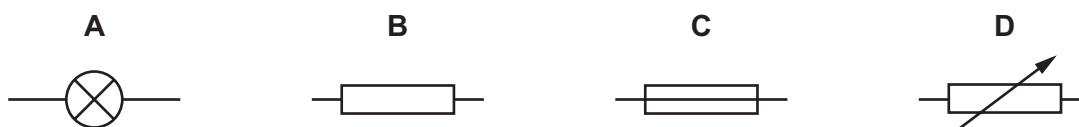
The ammeter measures the current in the heater.

The voltmeter measures the potential difference across the heater.

Which diagram shows this circuit?



- 39** What is the circuit symbol for a component whose only purpose is to protect an electric circuit?



- 40** Which risk is increased by using electrical equipment in damp conditions rather than in dry conditions?

- A** the cable to the equipment overheating
- B** the equipment overheating
- C** the current becoming too low
- D** the person receiving an electric shock

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The Periodic Table of Elements

Group																				
I	II											III	IV	V	VI	VII	VIII			
		1 H hydrogen 1																		
		Key atomic number atomic symbol name relative atomic mass																		
3 Li lithium 7	4 Be beryllium 9													5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19		
11 Na sodium 23	12 Mg magnesium 24													13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40	
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84			
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131			
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —			
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	113 Nh nihonium —	114 Fl flerovium —	115 Mc moscovium —	116 Lv livermorium —	117 Ts tennessine —	118 Og oganeson —			

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).