

Cambridge IGCSE™

CO-ORDINATED SCIENCES (DOUBLE AWARD)**0654/42**

Paper 4 Extended Theory

February/March 2024

MARK SCHEME

Maximum Mark: 120

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the February/March 2024 series for most Cambridge IGCSE, Cambridge International A and AS Level components, and some Cambridge O Level components.

This document consists of **11** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- 1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
- 2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
- 3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance
 For questions that require ***n*** responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
 - Any response marked *ignore* in the mark scheme should not count towards ***n***.
 - Incorrect responses should not be awarded credit but will still count towards ***n***.
 - Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
 - Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

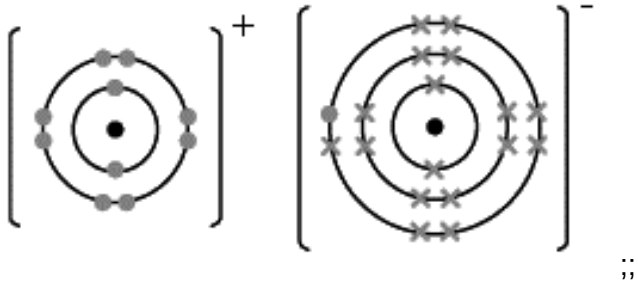
Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Question	Answer	Marks
1(a)	B ; E ; D ;	3
1(b)(i)	unpaired ;	1
1(b)(ii)	flagellum ; movement / locomotion ;	2
1(c)	egg (cell) ; red blood (cell) ; ciliated (cell) ; white blood (cell) ;	4
1(d)	plasma / platelets ;	1

Question	Answer	Marks
2(a)(i)	potassium sodium lithium ;	2
2(a)(ii)	lilac / purple / pink ;	1
2(a)(iii)	$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$;	2
2(b)(i)		2

Question	Answer	Marks
2(b)(ii)	Any two from: lattice (structure) ; regular arrangement ; of alternating positive and negative ions ;	2
2(b)(iii)	idea that ions can move in molten sodium chloride / ions cannot move in solid sodium chloride ;	1

Question	Answer	Marks
3(a)(i)	2.0 (cm) ;	1
3(a)(ii)	use of data from graph OR use of $F = kx$ OR $5.0 / 10$; ($k =$) 0.5 (N / cm) ;	2
3(a)(iii)	limit of proportionality ;	1
3(b)	volume using displacement method / eureka can ; mass using, balance / scales ; density = mass / volume ;	3
3(c)(i)	amplitude labelled from peak or trough to equilibrium position ;	1
3(c)(ii)	perpendicular / at right angles / 90° ;	1

Question	Answer	Marks
4(a)(i)	2 ; 4 ;	2
4(a)(ii)	Dd ;	1
4(a)(iii)	25 (%) ;	1

Question	Answer	Marks
4(b)	marasmus ; kwashiorkor ;	2
4(c)	glucose ; amino acids ; glucose ;	3
4(d)	protease ;	1

Question	Answer	Marks
5(a)(i)	LHS – anode ; RHS – cathode ;	2
5(a)(ii)	oxygen ;	1
5(a)(iii)	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$;	2
5(a)(iv)	(reduction) gain of electrons ;	1
5(a)(v)	test – lighted splint ; result – (squeaky) pop ;	2
5(b)	M_r of $\text{H}_2 = 2$; moles of $\text{H}_2 = 6 \div 24 = 0.25$; mass of $\text{H}_2 = 0.25 \times 2 = 0.5\text{g}$;	3

Question	Answer	Marks
6(a)	(1245 days =) 1.1×10^8 (s) ; (d=) $2\pi r$ OR $2\pi \times 3.8 \times 10^{11}$ OR 2.4×10^{12} (m) ; (distance / time) OR $2.4 \times 10^{12} / 1.1 \times 10^8$ (= 22196 OR 22000 m / s);	3
6(b)(i)	${}^0_{-1}\beta$;	1
6(b)(ii)	50 ; billion years ;	2
6(b)(iii)	7 (%) ;	1

Question	Answer	Marks
7(a)(i)	31 (bubbles / min) ;	1
7(a)(ii)	increases ; chemical ; chlorophyll ; carbohydrates ;	4
7(a)(iii)	fewer (bubbles of oxygen produced) ; carbon dioxide is a reactant of photosynthesis ;	2
7(b)	high temperature ; extremes of pH ;	2
7(c)(i)	phototropism ;	1
7(c)(ii)	auxin ;	1

Question	Answer	Marks
8(a)(i)	hydrogen ;	1
8(a)(ii)	temperature – 450 °C ; pressure – 200 atmospheres ;	2
8(a)(iii)	catalyst ; to speed up the (rate of) reaction / lower activation energy ;	2
8(b)	relative molecular mass of $\text{Fe}_2\text{O}_3 = 160$; $\frac{112 \times 400}{160} = 280$;	2
8(c)(i)	thermally decomposes ; calcium oxide. ; acidic / silica / sand ; slag / calcium silicate. ;	4
8(c)(ii)	(CaO is a basic oxide because) Ca / calcium is a metal (SiO_2 is an acidic oxide because) Si is a non-metal ;	1

Question	Answer	Marks
9(a)(i)	dull black (cube) absorbs (thermal / infrared) radiation better than dull white (cube) ;	1
9(a)(ii)	shiny black (cube) reflects radiation (rather than absorbs as the dull surface of black cube does) ;	1
9(a)(iii)	use of 3.0×10^8 (m / s) ; $(f =) v / \lambda / 3.0 \times 10^8 / 0.75 \times 10^{-3}$; $(f =) 4.0 \times 10^{11}$ (Hz) ;	3

Question	Answer	Marks
9(a)(iv)	vibrations (of ions / atoms) ; passes from one ion / atom to the next ; also by electrons moving through metal ;	3
9(b)(i)	two different metal wires joined at <u>two</u> junctions at different temperatures ;	1
9(b)(ii)	electromotive force and potential difference ticked ; electromotive force and potential difference ticked ; electromotive force only ticked ;	3

Question	Answer	Marks										
10(a)(i)	percentage of <u>energy</u> used from anaerobic respiration increases and the percentage of <u>energy</u> used from aerobic respiration decreases ;	1										
10(a)(ii)	3.5 (arbitrary units) ;	1										
10(b)	anaerobic respiration produces lactic acid ; (build-up of lactic acid) causes oxygen debt ; anaerobic respiration releases less energy <u>per glucose molecule</u> (than aerobic respiration) ;	3										
10(c)(i)	<table><tr><td>allows the passage of nervous impulses to body tissues</td><td></td></tr><tr><td>allows higher pressure of blood to the body tissues</td><td>✓</td></tr><tr><td>allows higher pressure of blood to the lungs</td><td></td></tr><tr><td>prevents diffusion of substances from the blood</td><td></td></tr><tr><td>separates oxygenated and deoxygenated blood</td><td>✓</td></tr></table> ;;	allows the passage of nervous impulses to body tissues		allows higher pressure of blood to the body tissues	✓	allows higher pressure of blood to the lungs		prevents diffusion of substances from the blood		separates oxygenated and deoxygenated blood	✓	2
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10(c)(ii)	contraction of (heart) muscle / atria / ventricle ;	1										

Question	Answer	Marks
11(a)	saturated ; single;	2
11(b)(i)	idea that the higher the number of carbon atoms the lower the energy given out / ORA ✓	1
11(b)(ii)	exothermic ;	1
11(c)	C ₁₄ H ₂₈ ;	1
11(d)	fuel oil / kerosene / naphtha ; idea that the supply is greater than the demand ;	2

Question	Answer	Marks
12(a)	(W =) $mgh / 1200 \times 10 \times 15$; (W =) 180000 (J) ;	2
12(b)(i)	correct shape of field ; correct direction indicated on at least one field line ;	2
12(b)(ii)	(Q =) $It / 50 \times 30$; (Q =) 1500 ; C / coulombs ;	3
12(b)(iii)	(V =) $IR / 50 \times 5.0$; (V =) 250 (V) ; (P =) $IV / 50 \times 250$; (P =) 12500 (W) ;	4
12(c)	can be switch off / on;	1