

CAMBRIDGE INTERNATIONAL EXAMINATIONS

Cambridge International Advanced Subsidiary and Advanced Level

MARK SCHEME for the March 2016 series

9701 CHEMISTRY

9701/22

Paper 2 (AS Structured Questions), maximum raw mark 60

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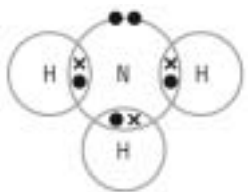
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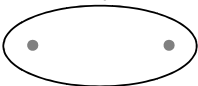
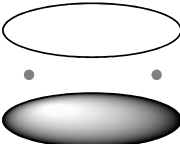
Page 2	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark	Total
1 (a) (i)	greater <u>attractive</u> force OR greater force <u>between nucleus and (outer) electrons</u> proton number / atomic number / nuclear charge increases across period AND electrons occupy same shell / shielding roughly constant	[1] [1]	[2]
(ii)	sulfur's electron removed from full (3p) <u>orbital</u> OR sulfur has two electrons in the same orbital electron–electron repulsion (reduces energy required)	[1] [1]	[2]
(iii)	sodium has mobile / free electrons / electrons free (to move throughout the structure) phosphorus is simple / covalent / molecular	[1] [1]	[2]
(iv)	magnesium has <u>two</u> free / delocalised / outer / valence electrons per atom OR <u>more</u> free / delocalised / <u>outer</u> electrons than sodium	[1]	[1]
(b) (i)	A = Mg(NO ₃) ₂ B = H ₂ C = NO ₂ OR O ₂ D = O ₂ OR NO ₂	[1] [1] [1] [1]	[4]
(ii)	any Group I carbonate OR ammonium carbonate	[1]	[1]
			[12]
2 (a) (i)	$\frac{27.30}{1000} \times 0.020 = 5.46 \times 10^{-4} \text{ (mol)}$	[1]	[1]
(ii)	(i) $\times 6 = 3.28 \times 10^{-3} \text{ (mol)}$	[1]	[1]

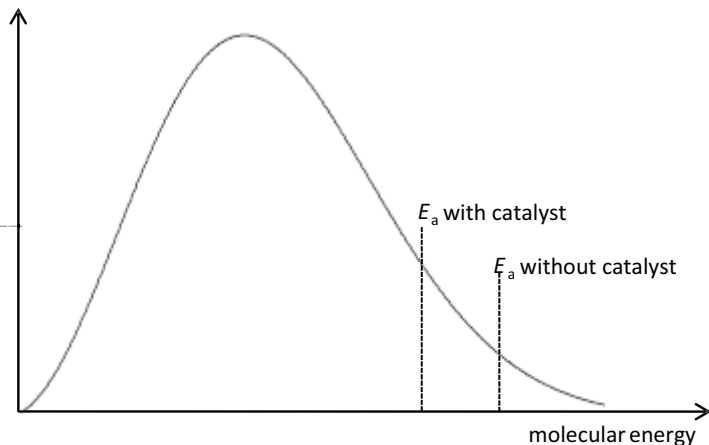
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Question	Answer	Mark	Total
(iii)	(ii) $\times \frac{250}{25.00} = 3.28 \times 10^{-2} \text{ (mol)}$	[1]	[1]
(iv)	$M_r \text{ of FeCO}_3 = 55.8 + 12.0 + 3(16.0) = 115.8$ (iii) $\times M_r(\text{FeCO}_3) = 3.79 \text{ g}$	[1] [1]	[2]
(v)	$\frac{\text{(iv)}}{5.00} \times 100\% = 75.9\%$	[1]	[1]
(b) (i)	$2\text{Fe}^{3+} + \text{Sn}^{2+} \rightarrow 2\text{Fe}^{2+} + \text{Sn}^{4+}$ species balancing	[1] [1]	[2]
(ii)	$\text{SnCl}_2(\text{aq}) + 2\text{HgCl}_2(\text{aq}) \rightarrow \text{SnCl}_4(\text{aq}) + \text{Hg}_2\text{Cl}_2(\text{s})$ SnCl_2 AND 2 state symbols	[1] [1]	[2]
			[10]
3 (a) (i)	 three bonding pairs lone pair AND octet shape = (trigonal) pyramidal	[1] [1] [1]	[3]

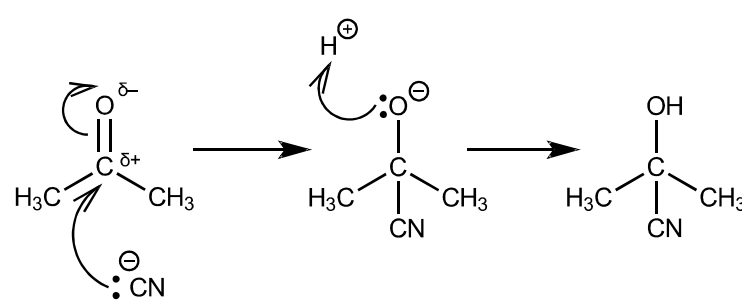
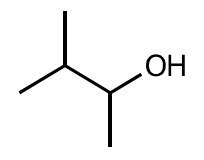
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Question	Answer	Mark	Total
(ii)	sigma(σ) bond  OR  pi(π) bond 	[1] [1]	[2]
(b) (i)	forward and backward reactions occurring <u>at same rate</u> OR <u>the rate of</u> forward and backward reactions are equal	[1]	[1]
(ii)	M1 = decreased yield of products/less products formed / ora M2 = <u>left</u>-hand side has fewer moles of gas OR equilibrium shifts to the <u>left</u>	[1] [1]	[2]

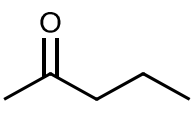
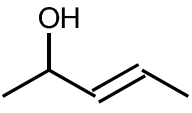
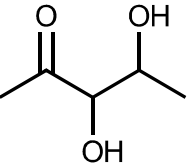
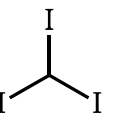
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Question	Answer	Mark	Total
(c)	 M1 = correct Boltzmann curve M2, M3 any 2 from: • line for both E_a values or statement in text that catalyst lowers E_a • (catalyst) increases proportion/number of molecules/particles with energy $\geq$ activation energy • so more frequent successful collisions	[1] [1] [1]	[3]

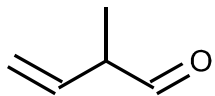
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Question	Answer	Mark	Total
(d) (i)	nucleophilic addition	[1]	[1]
(ii)	 correct dipole on carbonyl curly arrow from lone pair on CN⁻ AND from C=O to O correct intermediate curly arrow from lone pair on O⁻ to H⁺ correct product	[1] [1] [1] [1] [1]	[5]
			[17]
4 (a) (i)	<u>C₄H₁₀</u>	[1]	[1]
(ii)	<u>C₄H₉</u>	[1]	[1]
(iii)		[1]	[1]
(b)	$\text{C}_8\text{H}_{18} + 12\frac{1}{2}\text{O}_2 \rightarrow 8\text{CO}_2 + 9\text{H}_2\text{O}$	[1]	[1]
(c)	sulfur dioxide would be produced on combustion (which contributes to) <u>acid rain</u>	[1] [1]	[2]

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark	Total
(d)	M1 = H has more / greater / stronger van der Waals' / intermolecular forces than G / ora M2 = (because) H has more electrons (than G) M3 = J has hydrogen bonding (between molecules) M4 = strong(er) / great(er) forces require AND high / more energy to overcome	[1] [1] [1] [1]	[4]
(e)	NaOH(aq)	[1]	[1]
			[11]
5 (a) (i)	Q  R  S  T 	[1] [1] [1] [1]	[4]
(ii)	pent-3-en(e)-2-one OR 3-penten-2-one	[1]	[1]
(iii)	red / orange / yellow precipitate / solid	[1]	[1]

Page 8	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark	Total
(b)	<i>This question was discounted.</i> M1 = decolourises bromine / $1500\text{--}1600\text{ cm}^{-1}$ = alkene M2 = absorption at 1700 cm^{-1} is C=O AND (very) broad absorption at $2500\text{--}3000\text{ cm}^{-1}$ is O—H = carboxylic acid M3 = no cis-trans so terminal alkene OR chiral so contains a carbon atom with 4 different groups attached M4 = U is 	[1] [1] [1] [1]	[4]
			[10]