

MARK SCHEME for the May/June 2014 series

9701 CHEMISTRY

9701/22

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Page 2	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2014	9701	22

Question	Answers	Mark	Total
1 (a)	The (total) number of protons and neutrons (in the nucleus of an atom)	1	1
(b) (i)	Mass of an atom(s) or isotope relative to $\frac{1}{12}$ (the mass) of (an atom of) carbon–12 OR relative to carbon–12 which is (exactly) 12 (units) allow a correct expression	1 1	2
(ii)	^{79}Br ^{81}Br $78.92x$ $80.92(100-x)$ where x = % abundance of ^{79}Br so $\frac{78.92x + 80.92(100-x)}{100} = 79.9$ $x = 51$ hence $^{79}\text{Br} : ^{81}\text{Br} = 51 : 49$	1 1 1	3

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Question	Answers	Mark	Total
(c)	$\frac{A}{A_r} \quad \frac{Br}{79.9} = 1:3$ $\text{So } \frac{95.69/79.9}{4.31/A_r} = 3$ $A_r = \frac{3 \times 4.31 \times 79.9}{95.69} = 10.796 = 10.8 \text{ to 3 s.f.}$ 3 sig figs allow alternative correct methods	1 1 1	3
(d) (i)	Mg: bright/white light/flame OR white solid/smoke $\text{Mg} + \frac{1}{2} \text{O}_2 \rightarrow \text{MgO}$ allow correct multiples S: blue flame OR white/steamy fumes OR yellow solid disappears $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$ allow correct multiples	1 1 1 1	4
(ii)	$\text{Al}_2\text{O}_3 + 2\text{NaOH} + 7\text{H}_2\text{O} \rightarrow 2\text{NaAl}(\text{OH})_4(\text{H}_2\text{O})_2 \text{ OR}$ $\text{Al}_2\text{O}_3 + 2\text{NaOH} + 3\text{H}_2\text{O} \rightarrow 2\text{NaAl}(\text{OH})_4 \text{ OR}$ $\text{Al}_2\text{O}_3 + 2\text{NaOH} \rightarrow 2\text{NaAlO}_2 + \text{H}_2\text{O} \text{ OR}$ $\text{Al}_2\text{O}_3 + 2\text{OH}^- + 7\text{H}_2\text{O} \rightarrow 2[\text{Al}(\text{OH})_4(\text{H}_2\text{O})_2]^- \text{ OR}$ $\text{Al}_2\text{O}_3 + 2\text{OH}^- + 3\text{H}_2\text{O} \rightarrow 2[\text{Al}(\text{OH})_4]^- \text{ OR}$ $\text{Al}_2\text{O}_3 + 2\text{OH}^- \rightarrow 2\text{AlO}_2^- + \text{H}_2\text{O}$ $\text{Al}_2\text{O}_3 + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2\text{O}$ allow correct ionic equations	1 1	2

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Question	Answers	Mark	Total
(e)	shape of PCl_5 = (trigonal) bipyramid(al) bond angles in PCl_5 = 120° and 90°	1 1	2
			17
2 (a) (i)	(The $C_2O_4^{2-}$ ions) lose electrons owtte/ora	1	1
(ii)	$2MnO_4^-(aq) + 5C_2O_4^{2-}(aq) + 16H^+(aq) \rightarrow 2Mn^{2+}(aq) + 10CO_2(aq) + 8H_2O(l)$	1+1+1	3
(b) (i)	$\frac{20.0 \times 0.100}{1000} = 2(.00) \times 10^{-3} \text{ (mol)}$	1	1
(ii)	$MnO_4^- : C_2O_4^{2-} = 2 : 5$ so amount of $C_2O_4^{2-} = (5/2) \times 2.00 \times 10^{-3} = 5(.00) \times 10^{-3} \text{ (mol)}$ ecf from (b)(i)	1	1
(iii)	$5.00 \times 10^{-3} \times 250/25 = 0.05(0) \text{ (mol)}$ ecf from (b)(ii)	1	1
(iv)	amount = mass / M_r so M_r = mass / amount = $6.30/0.05 = 126$ ecf from (b)(iii)	1	1
(v)	$126 - 90 = 36$ $36/18 = 2.00$ $x = 2$ Ecf from (b)(iv) if suitable	1	1
			9

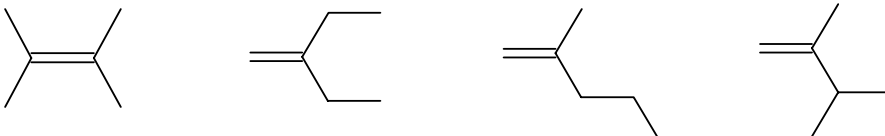
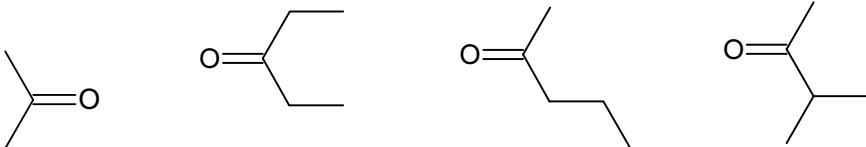
Page 5	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
3 (a) (i)	metallic bonding	1	3
	strength of attraction / metallic bonding increases (Na–Al) / more energy is needed to break 'bonds'	1	
	due to increasing cation charge / charge density / increasing number of delocalised electrons / decreasing ionic radius	1	
(ii)	van der Waals' (forces)	1	3
	are greatest / more in sulfur / relative magnitude of forces $S > P > Cl > Ar$	1	
	because sulfur has the greatest number of electrons / as no. of electrons (in the molecules) decreases	1	
(iii)	<u>Covalent</u> bond(s) broken OR (Si has a) giant covalent (structure)	1	1
(b) (i)	Nuclear charge (in Ar) greater (than Cl) AND same shielding owtte	1	1
(ii)	p subshell / orbital in Al at higher energy (than s subshell in Mg) ora OR p subshell / orbital more shielded ora	1	1
(iii)	repulsion due to electron pair (in same / p orbital)	1	1
			10

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Question	Answers	Mark	Total
4 (a)	physical: <u>fractional</u> distillation/fractionation	1	2
	chemical: crack(ing) (allow: reforming, isomerisation, thermal decomposition, desulfurisation)	1	
(b) (i)	Strong (C–C and C–H) bonds / high bond energies Non-polar / C and H have similar electronegativities	1 1	2
(ii)	109.5° AND 120° (117° – 122°)	1	1
(iii)	ethane = tetrahedral ethene = trigonal planar	1	1
(iv)	4 × σ /single bonds on Cs in ethane AND 3 × σ and 1 × π on Cs in ethene OR 2 × single and 1 × double on Cs in ethene allow from suitable labelled diagram	1	1
(c) (i)	$Cl_2 \rightarrow 2Cl\cdot$	1	5
	$C_2H_6 + Cl\cdot \rightarrow \cdot C_2H_5 + HCl$	1	
	$\cdot C_2H_5 + Cl_2 \rightarrow C_2H_5Cl + Cl\cdot$	1	
	$\cdot C_2H_5 + Cl\cdot \rightarrow C_2H_5Cl$	1	
	correct alternative terminations allowed		
	initiation, propagation, termination (correctly assigned)	1	
(ii)	$\cdot C_2H_5 + \cdot C_2H_5 \rightarrow C_4H_{10}$	1	1
			13

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
5 (a) (i)	decolourisation of bromine: P is an alkene / contains C=C / is unsaturated hot conc. manganate(VII): breaks C=C OR single product implies P is symmetrical OR single organic product implies terminal =CH₂ as methanal is oxidised to CO₂ 2,4-DNPH confirms >C=O / carbonyl / ketone in Q no reaction with Tollens': Q is not an aldehyde / is a ketone	1 1 1 1	4
(ii)	 (2,3-)dimethylbut-2-ene, 2-ethylbut(-1-)ene, 2-methylpent-1-ene, (2,3-)dimethylbut(-1-)ene	1 1	2
(iii)	 propan(-2-)one / acetone, pentan-3-one, pentan-2-one, 3-methylbutan(-2-)one ecf possible on (a)(ii)	1 1	2
(b)	any 3 of CH₃CH=C(CH₃)C₂H₅ CH₃CH=CHCH₂CH₂CH₃ CH₃CH=CHCH(CH₃)₂ C₂H₅CH=CHC₂H₅	1 1 1 1	max 3
			11