
CHEMISTRY

9701/42

Paper 4 A Level Structured Questions

May/June 2016

MARK SCHEME

Maximum Mark: 100

Published

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Question	Answer	Marks
1 (a) (i)	dative (covalent) <i>or</i> coordinate Hydrogen / H (bonding)	2
(ii)	octahedral	1
(iii)	$\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O} \rightarrow \text{Mg}(\text{NO}_3)_2 + 6\text{H}_2\text{O}$ $\text{Mg}(\text{NO}_3)_2 \rightarrow \text{MgO} + 2\text{NO}_2 + \frac{1}{2}\text{O}_2$ <i>any three of</i> (solid) dissolves / turns to liquid condensation on tube <u>white</u> solid (forms / remains) brown fumes (evolved) gas formed that relights a glowing splint	4
(iv)	M_r values: $\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O} = 256.3$ $\text{MgO} = 40.3$ or (loss in molar mass = $256.3 - 40.3 = 216$) percentage loss = $100 \times 216 / 256.3 = \mathbf{84.3 / 84.4\%}$	2
(b)	(cat)-ionic radius / ion size increases (down the group) less polarisation / distortion of nitrate ion / NO_3^-	2
(c)	$2\text{AgNO}_3 \rightarrow 2\text{Ag} + 2\text{NO}_2 + \text{O}_2$	1
		[Total: 12]
2 (a) (i)	(an acid that is) partially / incompletely ionised / dissociated	1
(b) (i)	$\text{p}K_a = -\log K_a$ or $K_a = 10^{-\text{p}K_a}$	1

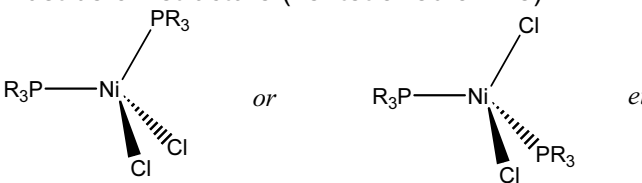
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Question	Answer	Marks
(ii)	ethanoic acid (1) is more acidic than propanoic acid (2) due to smaller electron-donating (R/ alkyl) group/ less electron-donating (R/ alkyl) group(s) 2-chloropropanoic acid (3) is more acidic than propanoic acid (2) due to electron-withdrawing/ electronegative (Cl/ chlorine) atom 2-chloropropanoic acid (3) is more acidic than 3-chloropropanoic acid (4) since the Cl/ chlorine/ electronegative atom is closer to the CO_2^-/ acid	3
(c) (i)	M1: voltmeter / V and salt bridge labelled M2: Cu and Cu^{2+} / CuSO_4 (any soluble Cu(II) salt) M3: H_2 (arrow in) and H^+ / HCl / H_2SO_4 / any mineral acid M4 Pt and one solution at 1 M / 1 mol dm^{-3} OR H_2 at 1 atm	4
(ii)	$E^\ominus_{\text{cell}} = 0.34 \text{ (V)}$ and $(\text{Cu}^{2+})/\text{Cu}$ is the positive electrode	1
d (i)	$K_a = 1.23 \times 10^{-5}$ $[\text{H}^+] = \sqrt{(K_a \cdot c)} = \sqrt{(1.23 \times 10^{-5} \times 0.1)} = 1.11 \times 10^{-3} \text{ mol dm}^{-3}$ pH = 3.0 (2.96) ecf from $[\text{H}^+]$	2

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Question	Answer	Marks
(ii)	$E = 0.0 + 0.059 \log(1.11 \times 10^{-3})$ OR $= -0.17(4)V$ so new $E_{\text{cell}} = 0.34 + 0.17 = \mathbf{0.51V}$ ecf from (d)(i)	2
		[Total: 14]
3 (a) (i)	$(\text{CH}_3)_2\text{CHCN}$	1
(ii)	reaction 1: NH_3 (in ethanol) under pressure (+ heat) or heat NH_3 in a sealed tube reaction 2: KCN/NaCN and heat/reflux (in ethanol) reaction 3: $\text{H}_2 + \text{Ni}$ or LiAlH_4	3
(b) (i)	$\text{CH}_3\text{CH}_2\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{CH}_2\text{NH}_3^+ (+) \text{OH}^-$	1
(ii)	ethylamine is more basic than ammonia... because of electron-donating (alkyl/ethyl/R) group (in ethylamine) which makes the <u>lone pair</u> (on N) more available for donation or the <u>lone pair</u> (on N) more available for a proton/ H^+	2
(c) (i)	A solution which resists/minimises/roughly maintains changes in <u>pH</u> when (small amounts of) H^+ or OH^- are added	1
(ii)	$\text{CH}_3\text{NH}_2 + \text{H}^+ \rightarrow \text{CH}_3\text{NH}_3^+$ $\text{CH}_3\text{NH}_3\text{Cl} + \text{OH}^- \rightarrow \text{CH}_3\text{NH}_2 + \text{H}_2\text{O} + \text{Cl}^-$	2
		[Total: 10]

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Question	Answer	Marks
4 (a) (i)	 (cis) (trans)	2
(ii)	cis is (more) polar due to both Cl^(δ-) on same side or cis is (more) polar as dipoles do not cancel / unsymmetrical or trans is non-polar as it is bond dipoles cancel	1
(iii)	(This can only be <i>cis</i>) its mirror image is the same / superimposable or the distance between two coordinating nitrogens / oxygens is too small to bond <i>trans</i> or difficult for the NH₂ and O to change places (since 5-membered rings can only bridge adjacent positions)	1
(b) (i)	It's not square planar or it's tetrahedral	1
(ii)	must be 3D structure (i.e. tetrahedral-like)  <i>or</i> <i>etc</i>	1
		[Total: 6]

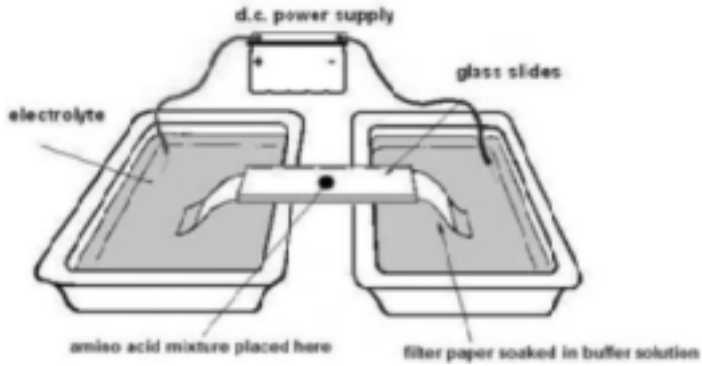
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Question	Answer	Marks
5 (a) (i)	$K_{\text{stab}} = \frac{[\text{Cd}(\text{CH}_3\text{NH}_2)_4^{2+}]}{[\text{Cd}^{2+}] [\text{CH}_3\text{NH}_2]^4}$ units: mol⁻⁴ dm¹²	2
(ii)	$\text{Cd}^{2+} + 4\text{CH}_3\text{NH}_2 \rightleftharpoons [\text{Cd}(\text{CH}_3\text{NH}_2)_4]^{2+}$ at start: 1×10^{-4} 0 at eqm: 1×10^{-7} y $1 \times 10^{-4} - 1 \times 10^{-7}$ or 9.99×10^{-5} or 1.0×10^{-4} $9.99 \times 10^{-5} / (\text{y}^4 \times 10^{-7}) = 3.6 \times 10^6$ and $\text{y} = \sqrt[4]{(9.99 \times 10^{-5}) / (1 \times 10^{-7} \times 3.6 \times 10^6)} = \mathbf{0.129/0.13}$	2
(b) (i)	(each complex is formed by) making (4 ×) N-Cd bonds and breaking (6 ×) O-Cd bonds or same types of/similar bonds forming/breaking or same number of bonds forming/breaking	1
(ii)	$\Delta S = (\Delta H - \Delta G) / T = (60.7 - 56.5) \times 1000 / 298 = (+)\mathbf{14}/(+)\mathbf{14.1}$	1
(iii)	fewer moles (of solutes) are forming (one mole of) the complex (so less loss of disorder) or one en displaces two H ₂ O whereas one CH ₃ NH ₂ only displaces one H ₂ O	1
(iv)	The [Cd(H ₂ NCH ₂ CH ₂ NH ₂) ₂] ²⁺ / equilibrium 2 complex (is more stable) because: <i>either</i> K _{stab} is greater or ΔG° is more negative.	1
		[Total: 8]

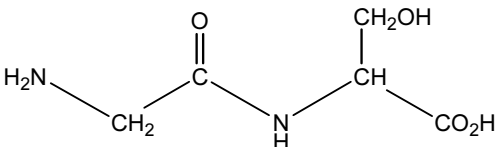
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Question	Answer	Marks
6 (a)	<i>essential mark</i> M1 the reactants/substrate has a shape complementary/specific to <u>active site</u> – can be awarded from a labelled diagram as below or diagrams showing this specificity clearly <i>any two of</i> M2: reactants/substrate binds to/fits into the <u>active site</u> of the enzyme M3: (Interaction with site) causes a specific bond to be weakened, (which breaks) or lowers activation energy M4: forms an E-S complex M5: products released from enzyme/active site labelled diagrams 	3
(b) (i)	δ 26 is CH₃-CO δ 52 is CH₃-O δ 169 is CH₃CO δ 167 is phenyl-CO <u>Phenyl ethanoate</u> is B <u>methyl benzoate</u> is A M1 = any two correct δ linked to phenylethanoate/methyl benzoate M2 = the rest correct	2
(ii)	heat with H₃O⁺ (to hydrolyse the ester) then add Br₂(aq)/bromine water decolourises/gives white ppt. (with phenol from B)	3
		[Total: 8]

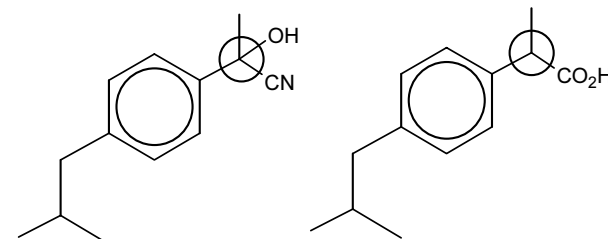
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Question	Answer	Marks
7 (a) (i)	<i>labelled with</i> M1: <u>DC</u> power supply + and – / battery / cell / + and – sign (on cell / electrodes) with a complete circuit M2: buffer solution / electrolyte labelled M3: (amino acid) mixture / x on (filter) paper / gel / agarose  The diagram shows an electrophoresis setup. A d.c. power supply is connected to two electrodes in two separate trays. The trays contain an electrolyte. A glass slide is placed across the trays, holding a strip of filter paper soaked in buffer solution. An amino acid mixture is placed on the filter paper between the two trays.	3
(ii)	direction of movement related to charge (of amino acids) distance travelled depends on charge / M_r (of amino acids)	2
(b) (i)	Asp + Val: pH 12 because Asp will be $-\text{CH}_2\text{COO}^-$ (R-group) moves further (to positive electrode than Val) or pH 12 Asp more negative so moves further (to positive electrode) or pH 12 because Asp has a charge of 2– but Val has a charge of 1– or best at pH 7 because Asp will be negatively charged (anionic) but Val neutral	1
(ii)	Lys + Ser: pH 2 because Lys will be $(\text{CH}_2)_4\text{NH}_3^+$ (R-group) moves further (to negative electrode than Ser) or pH 2 Lys more positive so moves further (to negative electrode) or pH 2 because Lys has a charge of 2+ and Ser has a charge of 1+ or pH 7 because Lys is positively charged (cationic) but Ser neutral / zwitterionic	1

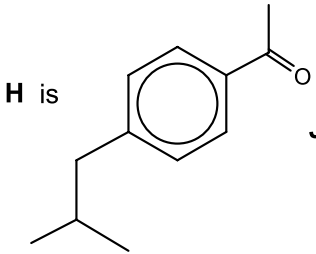
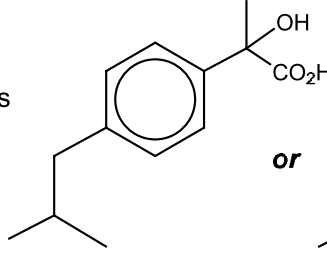
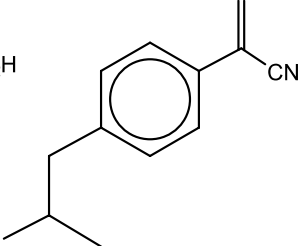
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(iii)	Tyr + Phe: pH 12 because Tyr will be $\text{C}_6\text{H}_5\text{CH}_2\text{O}^-$ (R-group) moves further / more / faster (to positive electrode than Phe) or pH12 because Tyr has a charge of 2– but Phe has a charge of 1–	1
(c) (i)	 M1: for –CONH– as shown above M2: for rest of molecule and correct connectivity of the bonds	2
(ii)	<i>from the IR spectrum</i> • E is O-H or N-H (allow NH_2) • F is C=O • G is C-O	2
		[Total: 12]
8 (a)	M1: solubility increases (down the group) M2: because lattice energy decreases faster than does ΔH_{hyd} M3: ΔH_{sol} / enthalpy of solution becomes more exothermic / less endothermic	3
(b) (i)	Should be the same / similar (enthalpy change), as (both acids) are fully ionised / strong acids	1

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Question	Answer	Marks
(ii)	$\text{Ca(s)} + 2\text{H}^+(\text{aq}) \longrightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2(\text{g})$  gas phase ions: $\text{Ca}^{2+}(\text{g}) + 2\text{H}^+(\text{g})$ $x = \Delta H_{\text{at}}(\text{Ca}) + \text{IE}(1) + \text{IE}(2) - 2\Delta H_{\text{hyd}}(\text{H}^+) + \Delta H_{\text{hyd}}(\text{Ca}^{2+}) - 2\text{IE}(\text{H}) - E(\text{H}-\text{H})$ $x = 178 + 590 + 1150 + 2(1090) - 1576 - 2(1310) - 436$ $x = -534 \text{ kJ mol}^{-1}$	4
(c)	$\text{CH}_3\text{CO}_2\text{H}$ is incompletely ionised / weak acid / weaker acid enthalpy change of ionisation (of CH_3COOH) is $+2 \text{ kJ mol}^{-1}$ or energy needed to ionise / dissociate (CH_3COOH)	2
		[Total: 10]
9 (a)		1

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Question	Answer	Marks
(b)	H is  J is  J1 <i>or</i>  J2	2
(c)	step 1: $(\text{CH}_3)_2\text{CHCH}_2\text{Cl} + \text{AlCl}_3$ (+ heat) step 2: $\text{CH}_3\text{COCl} + \text{AlCl}_3$ (+ heat) step 3: $\text{HCN} + \text{NaCN}$ <i>or</i> $\text{HCN} + \text{base}$ <i>or</i> $\text{HCN} + \text{CN}^-$ (steps 4 and 5 could be reversed on J) If J1 step 4 then step 5 J2 step 5 then step 4 step 4: H_3O^+ + heat/aqueous HCl + heat step 5: conc H_2SO_4 + heat/ conc H_3PO_4 + heat <i>or</i> Al_2O_3 + heat step 6: $\text{H}_2 + \text{Ni}$ (+ heat)	6
(d)	step 1: electrophilic substitution <i>or</i> alkylation step 6: reduction / hydrogenation / addition	2
		[Total: 11]

